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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds $\text{Cl}^- \text{aq}$

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{ 2Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4^+ \text{aq} + 2\text{F}^- \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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