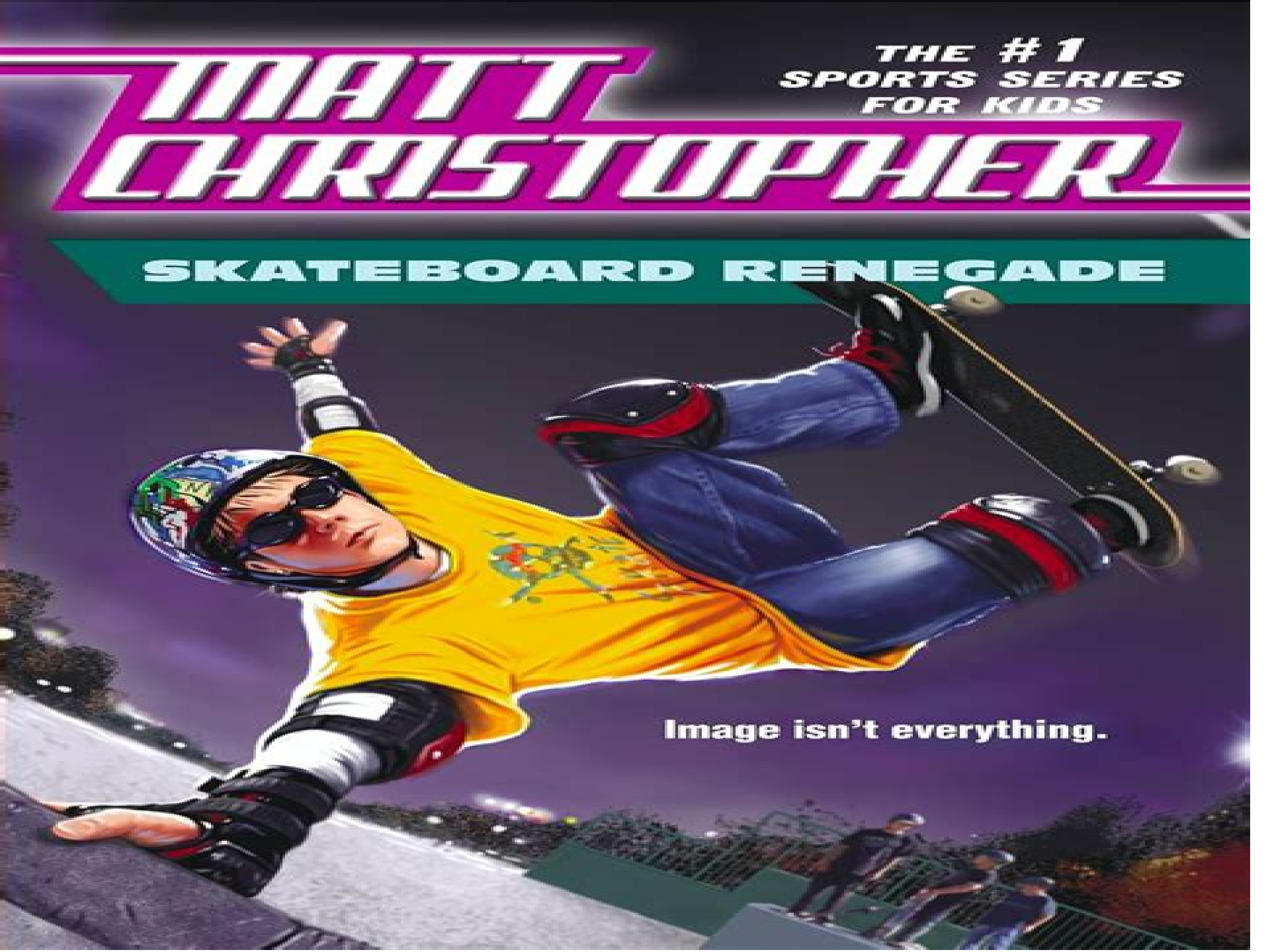




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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27.0 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} = 1.11 \text{ mol } \text{Al}_2\text{O}_3$   $1.11 \text{ mol } \text{Al}_2\text{O}_3 \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 113.2 \text{ g } \text{Al}_2\text{O}_3$  c determine the percent yield

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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## **limiting reagents and percentage yield worksheet answers** - Sep 04 2022

web limiting reagents and percentage yield worksheet 1 consider the reaction i<sub>2</sub>o<sub>5</sub> g 5 co g 5 co<sub>2</sub> g i<sub>2</sub> g a 80 0 grams of

iodine v oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  to determine the mass of iodine  $\text{I}_2$  which could be produced.  
 $80 \text{ g } \text{I}_2\text{O}_5 \times \frac{1 \text{ mol } \text{I}_2\text{O}_5}{1 \text{ mol } \text{I}_2} \times \frac{1}{333.8 \text{ g } \text{I}_2\text{O}_5} \times 28 \text{ g } \text{CO} \times \frac{1 \text{ mol } \text{CO}}{28 \text{ g } \text{CO}} = 1 \text{ mol } \text{I}_2$

### 7.3 limiting reactant and percent yield problems - Jun 13 2023

web Jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage.  
 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$  label 3.7.3