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the order of the reaction with respect to  $A$  the value of  $n$  is not related to the reaction stoichiometry and

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sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the  
first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2N_2O_5 \rightarrow 4NO_2 + O_2$  has been found  
to be directly proportional to the concentration of  $N_2O_5$  text rate  $k$   $n$

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide ki sodium thiosulfate na<sub>2</sub>s<sub>2</sub>o<sub>3</sub> starch and hydrogen peroxide h<sub>2</sub>o<sub>2</sub> were used in order to calculate the rate of reaction

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