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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN}_2^- + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{H}_2\text{O} \rightarrow 2\text{AuCN}_2^- + 2\text{H}_2\text{O} + \text{O}_2$ multiplied by a factor of 2 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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