

Author of the Newbery Medal-winning JULIE OF THE WOLVES

JEAN CRAIGHEAD GEORGE

AN ECO MYSTERY

The Fire Bug Connection



The Fire Bug Connection An Ecological Mystery

Lynn Plourde, Paul Knowles



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions

gives the net ionic equation which shows only those species that participate in the chemical reaction

$$2\text{Ag}^+(aq) + \text{Cr}_2\text{O}_7^{2-}(aq) \rightarrow 2\text{Cr}^{3+}(aq) + 7\text{H}_2\text{O}(l) + 2\text{Ag}^+(aq)$$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is

$$2\text{Na}^+(aq) + 2\text{Cl}^-(aq) + \text{Cu}^{2+}(aq) + \text{SO}_4^{2-}(aq) \rightarrow 2\text{Na}^+(aq) + \text{SO}_4^{2-}(aq) + \text{CuCl}_2(s)$$

the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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$$\text{H}^+(aq) + \text{Cl}^-(aq) + \text{Na}^+(aq) + \text{HCO}_3^-(aq) \rightarrow \text{Na}^+(aq) + \text{Cl}^-(aq) + \text{H}_2\text{O}(l) + \text{CO}_2(g)$$

the spectator ions are Cl^- and Na^+

net ionic equation

$$\text{H}^+(aq) + \text{HCO}_3^-(aq) \rightarrow \text{H}_2\text{O}(l) + \text{CO}_2(g)$$

here is a video on spectator ions

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web the six Na^+ ions and the six Cl^- ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation

$$3\text{Ba}^{2+}(aq) + 2\text{PO}_4^{3-}(aq) \rightarrow \text{Ba}_3(\text{PO}_4)_2(s)$$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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$$\text{AgNO}_3(aq) + \text{KCl}(aq) \rightarrow \text{AgCl}(s) + \text{KNO}_3(aq)$$

total ionic

$$\text{Ag}^+(aq) + \text{NO}_3^-(aq) + \text{K}^+(aq) + \text{Cl}^-(aq) \rightarrow \text{AgCl}(s) + \text{K}^+(aq) + \text{NO}_3^-(aq)$$

net ionic

$$\text{Ag}^+(aq) + \text{Cl}^-(aq) \rightarrow \text{AgCl}(s)$$

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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