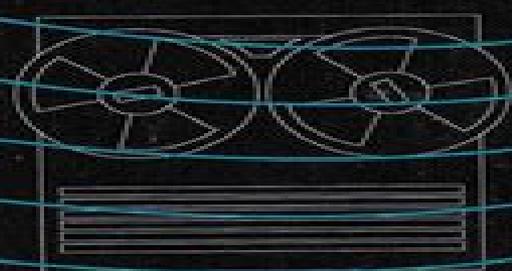
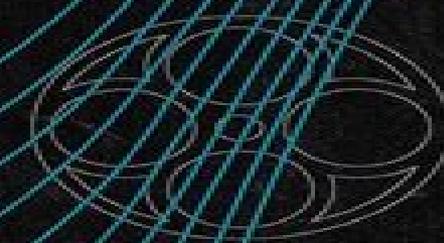
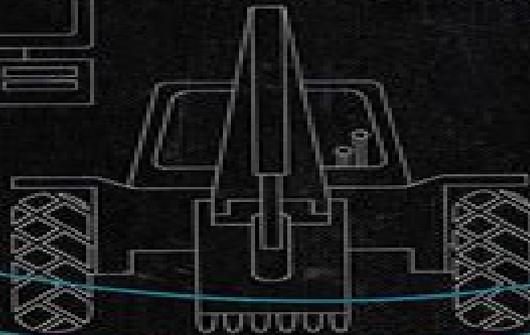


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web nov 23 2017 maharashtra x physics suppose the masses of calorimeter the water in it and the hot object made up of copper which is put in the calorimeter are the same the initial temperature of the calorimeter and water is 30 °C and that of hot object is 60 °C the specific heat of copper and water are 0.09 cal/g °C and 1 cal/g °C respectively

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web oct 28 2019 solution use this equation  $q = \text{specific heat} \times m \times \Delta t$  where  $q$  is heat flow  $m$  is mass in grams and  $\Delta t$  is the temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ }^\circ\text{C} = 1250 \text{ J}$   $q_{\text{water}} = 550 \text{ J}$   $\Delta h_{\text{water}} = 550 \text{ J}$

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web jul 12 2023 this concept lies at the heart of all calorimetry problems and calculations because the heat released or absorbed at constant pressure is equal to  $\Delta h$  the relationship between heat and  $\Delta h_{rxn}$  is  $\Delta h_{rxn} = q_{rxn} / n_{mcs} \Delta t$

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25 0 c and heat it to 145 0 c what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have 18 7 grams of an unknown liquid chemical at a temperature of 35 0 c

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web sample problem calorimetry and enthalpy changes in an experiment 25 0 ml of 1 00 m hcl at 25 0 c is added to 25 0 ml of 1 00 m naoh at 25 0 c in a foam cup calorimeter a reaction occurs and the temperature rises to 32 0 c calculate the enthalpy change in kj for this reaction assume the densities of the solutions are 1 00 g ml and that

6 4 calorimetry chemistry libretexts - Jan 31 2023

web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal m and a low temperature substance such as cool water w answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to  $t_1$  than to  $t_2$  a the final temperature is closer to  $t_1$  than to  $t_2$  the final temperature is exactly halfway between  $t_1$  and  $t_2$  b the final temperature is exactly halfway between  $t_1$  and  $t_2$  the final temperature is closer to  $t_2$  than to  $t_1$  c

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web calorimetry worksheet 1 if 0.315 moles of hexane  $C_6H_{14}$  is combusted in a bomb calorimeter containing 5.65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 5.54 c the specific heat capacity of water is 4.184 J/g C

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web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from ice at 34°C to steam at 138°C a sample of water absorbs  $1.41 \times 10^5$  J find the mass of the sample answers 1 2 68 x 10<sup>4</sup> J 2 6 23 x 10<sup>4</sup> J 3 1 71 x 10<sup>4</sup> J 4 1 11 x 10<sup>5</sup> J 5 2 11 x 10<sup>5</sup> J 6

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