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Snarling Citizen Essays

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S $\times$ 18 g H₂O / 1 mole H₂O = 149.4 g H₂O is your theoretical yield so percent yield = 137.1 g H₂O actual yield / 149.4 g H₂O $\times$ 100 = 91.77% problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = mass of product obtained / calculated mass of the product expected $\times$ 100 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mole Al}_2\text{O}_3}{4 \text{ mole Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 160.0 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 = \frac{126.0 \text{ g}}{160.0 \text{ g}} \times 100 = 78.75\%$$

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is
 reacted with 8.6 g of H_2

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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4 + \text{NaNO}_3$ which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 x 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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