

# SLAVERY & *the* ROMANTIC IMAGINATION



William L. G. S.

# Slavery And The Romantic Imagination

**Puja Mehta**



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web but it was not until june 9 1894 that texas had a major discovery this occurred in the drilling of a water well for the city of corsicana oil caused that well to be abandoned but a company formed in 1895 drilled several producing oil wells the first well equipped refinery in texas was built in corsicana in 1898 and this plant which

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web limiting reactant theoretical and percentage yields key 4  $\text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  
$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$
  
8.3 mole  $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  
$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  
$$\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mole}} = 0.666 \text{ moles } \text{O}_2$$

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web a write a balanced equation for the reaction  
$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$
  
b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   
$$160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3/\text{mol} = 302.3 \text{ g Al}_2\text{O}_3$$
  
c determine the percent yield  
$$\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$$

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web 7.NA.2.C.2.O.4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of  $\text{H}_2$  are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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