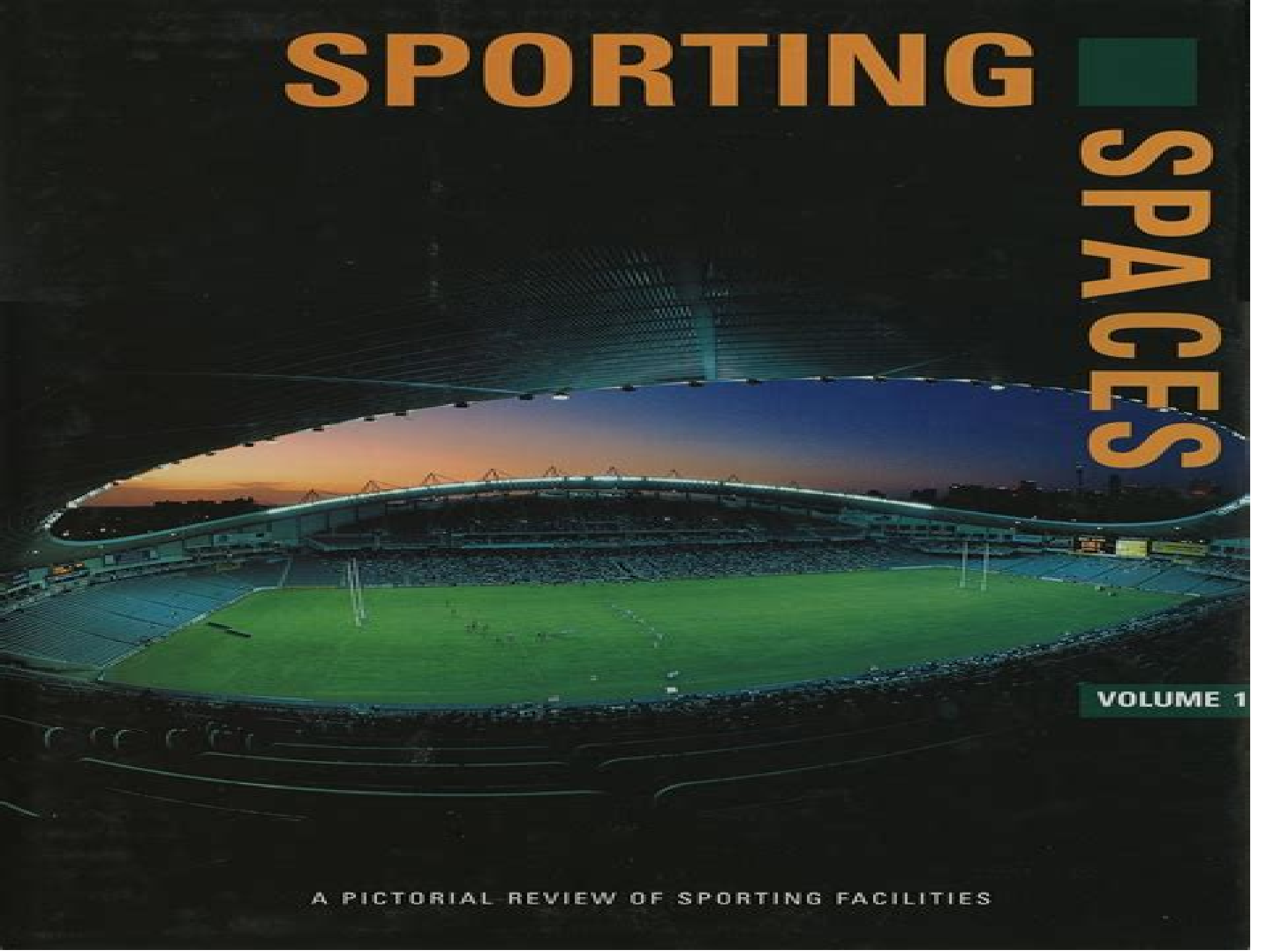


SPORTING

An aerial night photograph of a large sports stadium. The stadium has a distinctive curved, illuminated roof structure. The field is green and lit up, with goalposts visible. The stands are filled with spectators. The sky is dark, and the city lights are visible in the background.

SPACES

VOLUME 1

A PICTORIAL REVIEW OF SPORTING FACILITIES

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} \times 18 \text{ g } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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you have 0.216232 grams/mole $\times$ 21.3 grams = 0.66 moles $\times$ 32 g/mol

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calculated mass of the product expected $\times$ 100 in this reaction a yield of 80 is anticipated

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g Al_2O_3 $\times$ 1 mol Al_2O_3 $\times$ 101.964 g Al_2O_3 / 302.3 g Al_2O_3 = 26.982 g Al_2O_3 c determine the percent yield

yield actual $\times$ 100 260.0 g $\times$ 100 / 86.01 theoretical 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be

weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of H₂ are left over 13

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is

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