

Name	Formula	Explanation
Area of a rectangle	$A = l \times w$ or $A = lw$	where $l$ is the length and $w$ is the width
Area of a triangle	$A = \frac{1}{2}b \times h$ or $A = \frac{1}{2}bh$	where $b$ is the base and $h$ is the perpendicular height
Area of a parallelogram	$A = b \times h$ or $A = bh$	where $b$ is the base and $h$ is the perpendicular height
Area of a trapezium	$A = \frac{a+b}{2} \times h$ or $A = \frac{(a+b)h}{2}$	where $a$ and $b$ are the parallel sides and $h$ is the perpendicular height
Circumference of a circle	$C = 2\pi r$ or $C = \pi D$	where $r$ is the radius and $D$ is the diameter
Area of a circle	$A = \pi r^2$	where $r$ is the radius
Volume of a prism	$V = \text{area of cross section} \times h$	where $h$ is the height
Volume of a cylinder	$V = \pi r^2 h$	where $r$ is the radius and $h$ is the height
*Volume of a pyramid	$V = \frac{1}{3} \text{ area of base} \times h$	where $h$ is the perpendicular height
Pythagoras' theorem	$a^2 + b^2 = c^2$	where $a$ , $b$ and $c$ are side lengths of a right angled triangle, and $c$ is the hypotenuse
Trigonometric ratios (SOH CAH TOA)	$\sin x = \frac{\text{opp}}{\text{hyp}}$ $\cos x = \frac{\text{adj}}{\text{hyp}}$ $\tan x = \frac{\text{opp}}{\text{adj}}$	
*Sine rule	$\frac{a}{\sin A} = \frac{b}{\sin B} = \frac{c}{\sin C}$	where $A B C$ are angles and $a b c$ are sides of a triangle 
*Cosine rule	$a^2 = b^2 + c^2 - 2bc \cos A$	
*Area of a triangle (trigonometry)	$A = \frac{1}{2}ab \sin C$	
*The quadratic formula (solutions to a quadratic equation)	$x = \frac{-b \pm \sqrt{(b^2 - 4ac)}}{2a}$	where $a$ , $b$ and $c$ are the coefficients in the quadratic equation $ax^2 + bx + c = 0$
Compound interest	$\text{Total} = P(1 + \frac{r}{100})^n$	where $P$ is the principal amount, $r$ is the interest rate over a given period and $n$ is number of times that the interest is compounded
Probability "and" rule	$P(A \text{ or } B) = P(A) + P(B) - P(A \text{ and } B)$	where $P(A)$ is the probability of outcome $A$ and $P(B)$ is the probability of outcome $B$
Probability "or" rule	$P(A \text{ and } B) = P(A \text{ given } B)P(B)$	

# The Ebentials Of Edexcel Gcse Maths

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*periodic trends straw lab periodic trends straw lab studocu* - Dec 31 2022

web periodic trends straw lab objective to create a visual representation of the following periodic trends atomic size ionization energy and electronegativity atomic radius the size of an atom measured in either nanometers nm angstroms Å

periodic trends guided inquiry activity chemical education - Jun 24 2022

web jan 28 2019 by mapping specific trends directly onto a blank periodic table the activity uses the periodic table as a foundational aspect of the major trends allowing for students to make connections between placement on the periodic table and a specific trend

**periodic trends practice periodic table khan academy** - Feb 01 2023

web you might need periodic table the first two ionization energies for beryllium are shown below be g be g e be g be 2 g e i 1 900 kJ mol i 2 1757 kJ mol which of the following identifies the most probable value for the third ionization energy for be and provides the best justification

*periodic trends national science teaching association* - Mar 02 2023

web the purpose of this lab is to introduce students to periodic trends it gives students an opportunity to explore the values of different properties for all elements in the periodic table to determine which properties of the elements follow a periodic trend

**chemistry periodic trends lab pdf slideshare** - Feb 18 2022



web 1 name date per periodic trends lab periodic law states that when the elements are arranged in order of increasing atomic number there is a periodic repetition of their chemical and physical properties elements within the same group have the most similar chemical and physical properties although not identical

**periodic trends gizmo for the gizmo that 2018 periodic trends answer** - Oct 09 2023

web periodic trends answer key vocabulary atomic radius electron affinity electron cloud energy level group ion ionization energy metal nonmetal nucleus period periodic trends picometer valence electron

**dry lab periodic trends answer key pdf scribd** - Mar 22 2022

web 1 rank the following elements by increasing atomic radius carbon aluminum oxygen potassium oxygen carbon aluminum potassium 2 rank the following elements by increasing electronegativity sulfur oxygen neon aluminum neon aluminum sulfur oxygen 3 why does fluorine have a higher ionization energy than iodine

**simulation periodictrends answerkey google docs** - Jun 05 2023

web simulation periodic trends answer key background in this investigation you will examine several periodic trends including atomic radius ionization energy and ionic radius you will

**lab 5 periodic trends graphing lab** - Apr 03 2023

web 1 how does atomic radius change as you move from left to right across a period explain your answer in terms of the number of protons in the nucleus 2 how does atomic radius change as you move from top to bottom in a group explain your answer in terms of the number of electron shells 3 on the representation of the periodic table below

dry lab periodic trends answer key name studocu - Aug 07 2023

web periodic trends worksheet directions use your notes to answer the following questions rank the following elements by increasing atomic radius carbon aluminum oxygen potassium oxygen carbon aluminum potassium rank the following elements by increasing electronegativity sulfur oxygen neon aluminum

**periodic trends lab flashcards quizlet** - Sep 08 2023

web caragabriel preview chemistry test oct 7 11 43 terms misterlominyo preview terms in this set 25 each element has 3 numbers listed under it which value represents the atomic radius the first number below the diagram what are the units for the atomic radius pm picometers 10 negative twelve meters

*10a periodic trends worksheet chemistry libretexts* - Jul 26 2022

web mar 13 2023 across the periodic table sizes of atoms show the following trends with many irregularities size increases down a group the outermost electrons are in successively more extensive orbitals as n increases

*04 periodic trends lab periodic trends and the properties* - Aug 27 2022

web similarities and differences among the elements give rise to so called periodic trends both across rows and within

columns of the periodic table recognizing periodic trends in the physical and chemical properties of the elements is key to understanding the full value of the periodic table concepts periodic table periodic trends

*solved periodic trends straw lab the purpose of this lab is chegg* - Sep 27 2022

web question periodic trends straw lab the purpose of this lab is to create a visual representation of trends on the periodic table you will need answer the following questions and write a formal lab report use the information below to help guide you through writing your lab report 1