

THE **AMERICAN** **FORUM**

IN PURSUIT OF THE
DREAM, 1948 - 1993

GI



HENRY A. J. RAMOS

THE AMERICAN GI FORUM

HENRY A. J. RAMOS

The American Gi Forum In Pursuit Of The Dream 19481983

J Dewey



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furthermore using the stoichiometric ratio between the naoh and the acid analyte 1 1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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 $\text{HCl (aq)} + \text{NaOH (aq)} \rightarrow \text{NaCl (aq)} + \text{H}_2\text{O (l)}$
one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH

reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \text{ aq}$ $2\text{NaOH} \text{ aq}$ $\text{Na}_2\text{C}_2\text{O}_4 \text{ 4H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction $\text{NaOH} \text{ aq}$ $\text{C}_6\text{H}_4\text{COOH} \text{ aq}$ $\text{C}_6\text{H}_4\text{COONa} \text{ aq}$ H_2O the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the

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