

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: ionic bond and covalent bond

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of an



Bond and Electron sharing



Bonding energy



There are two types of

Definition of electron



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds

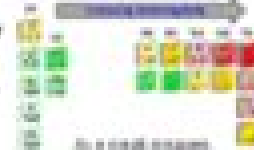
- Sigma bonds are formed by head-on (end-to-end) overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- One single bond is a sigma bond.
- One double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electron is shared equally (nonpolar bond). For example, the C-C and C-H bonds.

- If the difference in electronegativity is between 0.5 and 1.7, the bond is a polar covalent bond. For example, the C-O and C-N bonds.

- A difference of 1.7 or higher is ionic, so the electron is not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moment

The geometry of the molecule is important. Symmetrical molecules are nonpolar (even for dipole moment) are cancelled. A good example is carbon dioxide.



Cancel the vector as the other pointing point in the same way

Bond Length and Bond Strength

The longer the bond, the weaker the bond



Double bonds are stronger than single bonds and are shorter bonds

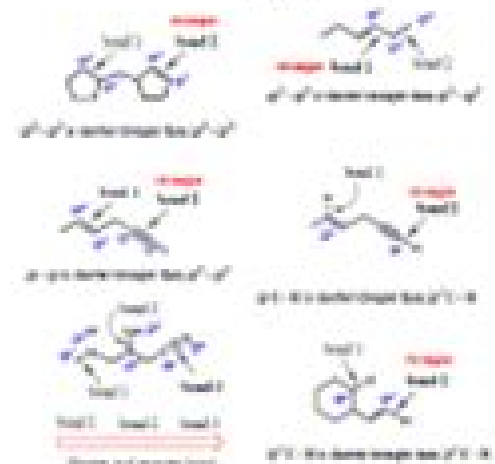


The longer the bond, the weaker the bond

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3 C-C	sp^2-sp^2 C-C	$sp-sp$ C-C
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	104 kcal/mol	110 kcal/mol

Higher s character, shorter and stronger bond

Example



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**Robert Thornton Morrison, Robert
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Table of Contents Study Guide For Organic Chemistry

1. Understanding the eBook Study Guide For Organic Chemistry
 - The Rise of Digital Reading Study Guide For Organic Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Organic Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Organic Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Organic Chemistry
 - Personalized Recommendations
 - Study Guide For Organic Chemistry User Reviews and Ratings

- Study Guide For Organic Chemistry and Bestseller Lists
- 5. Accessing Study Guide For Organic Chemistry Free and Paid eBooks
 - Study Guide For Organic Chemistry Public Domain eBooks
 - Study Guide For Organic Chemistry eBook Subscription Services
 - Study Guide For Organic Chemistry Budget-Friendly Options
- 6. Navigating Study Guide For Organic Chemistry eBook Formats
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 - Study Guide For Organic Chemistry Compatibility with Devices
 - Study Guide For Organic Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide For Organic Chemistry
 - Highlighting and Note-Taking Study Guide For Organic Chemistry
 - Interactive Elements Study Guide For Organic Chemistry
- 8. Staying Engaged with Study Guide For Organic Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Organic Chemistry
- 9. Balancing eBooks and Physical Books Study Guide For Organic Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Organic Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Organic Chemistry
 - Setting Reading Goals Study Guide For Organic Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Organic Chemistry
 - Fact-Checking eBook Content of Study Guide For Organic Chemistry
 - Distinguishing Credible Sources

-
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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