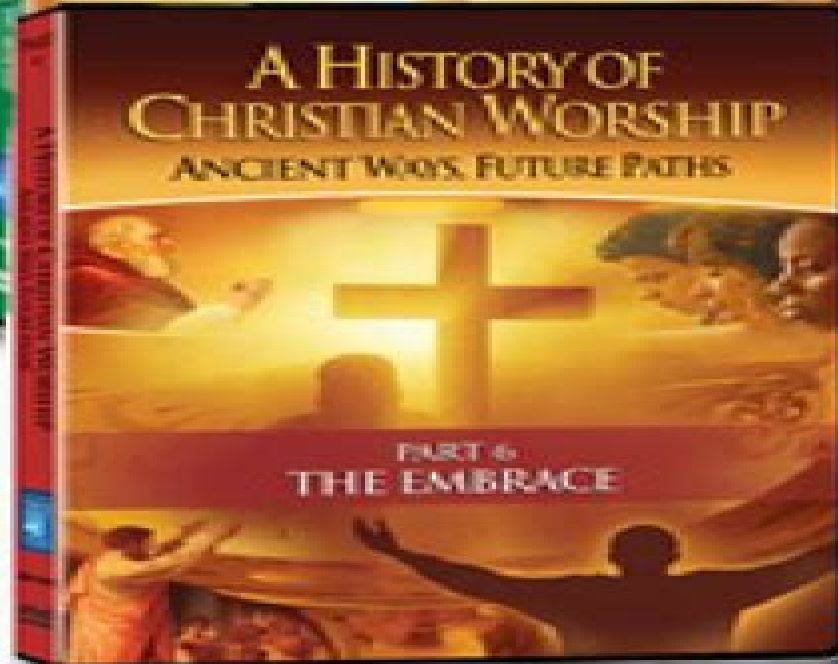
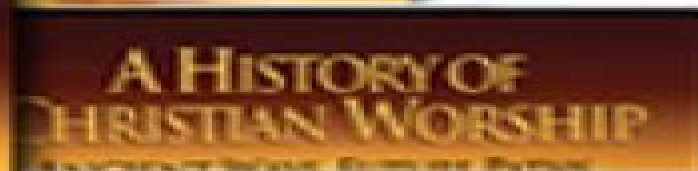
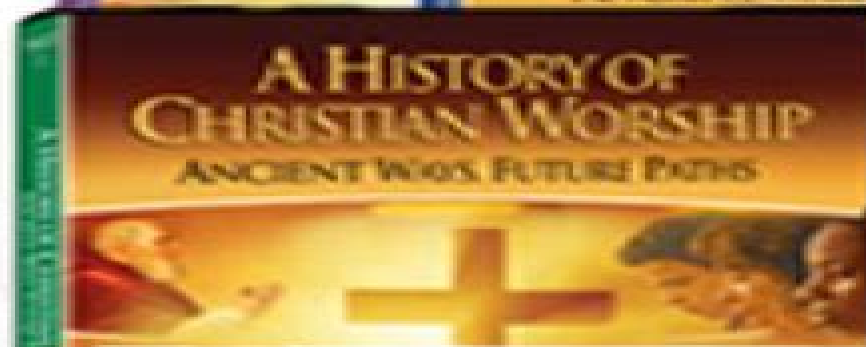
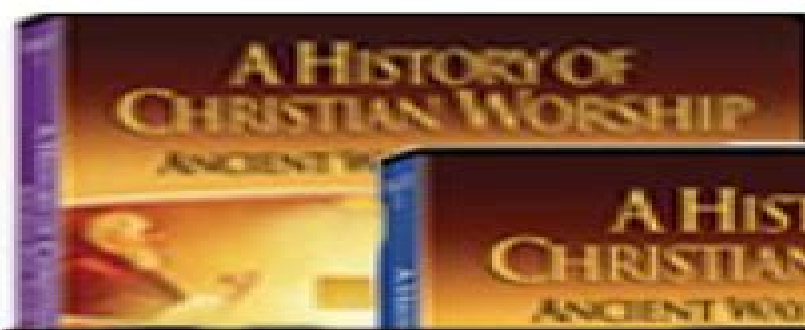




The Gospel And The Catholic Church Classics Series

Wolfgang Guggemos



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mentioned in ancient authors ... Together with an account of coins, weights and measures, etc Charles ANTHON
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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b Ba^{2+} c Br_2 d ca e Ga^{3+} f h 2g h

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what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + \text{I}^-$ 2 $\text{O}_5 + 5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation
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reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} oxidized reducing agent O_2 to O^{2-} reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+1} oxidized red ag
 S^0 to S^{2-} reduced ox ag $3\text{Cs} + \text{Br}_2$