

The Great Within



Christian D. Larson

The Great Within

Christian D Larson



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web a at atmospheric pressure helium can exist in all three phases as well as a supercritical fluid phase near absolute zero
 choice b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure b solid and
 gaseous helium never exist in equilibrium with each other at any temperature or pressure choice c

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 enthalpy change is path independent note with regard to choice b that enthalpy change equals internal energy change only if
 the volume of the system remains constant and thus no thermodynamic work is performed 5 b

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web 1 answer a explanation the value of ΔH is negative therefore the reaction is exothermic or heat releasing 2 answer a 1
 explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved 3 answer c explanation option c is
 incorrect because 566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes $\text{H}_2\text{O}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 6.01 \text{ kJ/mol}$ this is an endothermic process it requires 6.01 kJ to melt one mole of ice $\text{H}_2\text{O}(\text{s})$ the enthalpy value will change if the number of moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy ΔE of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings $\Delta E = q + w$
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