



# Small Fatigue Cracks

**SA Adler**



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yield 3co 7h 2 c 3 h 8 3h 2 o

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web a write a balanced equation for the reaction al + o<sub>2</sub> → al<sub>2</sub>O<sub>3</sub> 4 al + 3 o<sub>2</sub> → 2 al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of al<sub>2</sub>O<sub>3</sub> 160 0 g al x 1 mol al / x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al / 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual x 100 260 0 g / x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h<sub>2</sub> are left over 13

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yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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