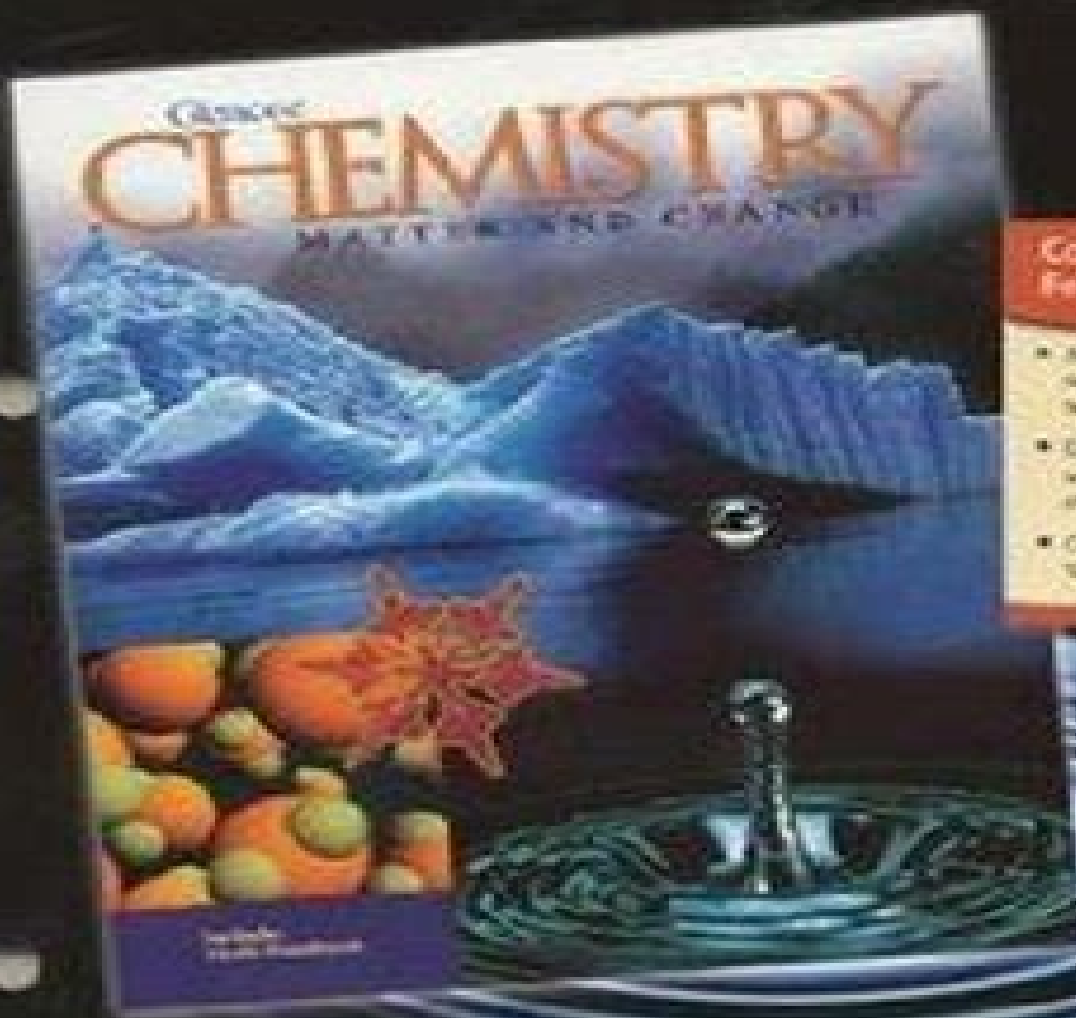


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**Raffaella Di Napoli**



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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of  $1s^2 2s^1$

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web the electron removed during the ionization of beryllium  $1s^2 2s^2$  is an s electron whereas the electron removed during the ionization of boron  $1s^2 2s^2 2p^1$  is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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