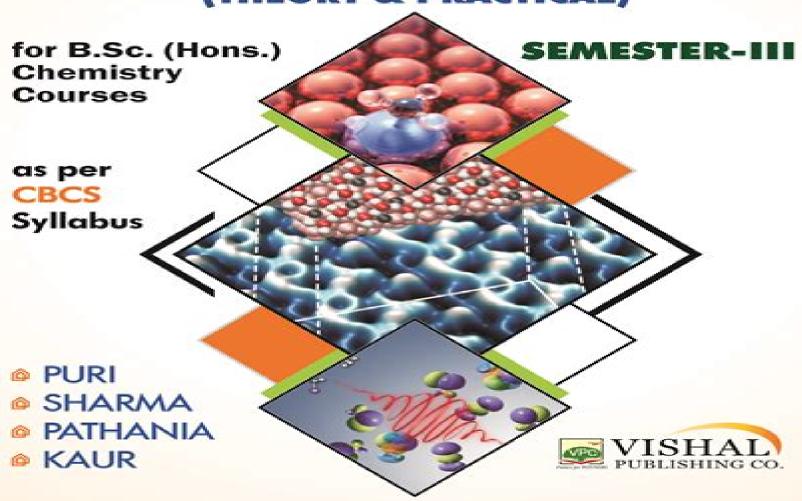
A Textbook of Physical Chemistry-III

(THEORY & PRACTICAL)



Textbook Of Physical Chemistry

Nk Vishnoi / Rj Shukla

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disciplines of Physical Chemistry Admittedly the text is loaded with some mathematics which is a prerequisite to thoroughly understand the topics presented here However the mathematics needed is explained at a really low level so that no additional mathematical textbook is needed A Textbook of Physical Chemistry - Volume 1 Mandeep Dalal, 2018-01-01 An advanced level textbook of physical chemistry for the graduate B Sc and postgraduate M Sc students of Indian and foreign universities This book is a part of four volume series entitled A Textbook of Physical Chemistry Volume I II III IV CONTENTS Chapter 1 Quantum Mechanics I Postulates of quantum mechanics Derivation of Schrodinger wave equation Max Born interpretation of wave functions The Heisenberg's uncertainty principle Quantum mechanical operators and their commutation relations Hermitian operators elementary ideas quantum mechanical operator for linear momentum angular momentum and energy as Hermition operator The average value of the square of Hermitian operators Commuting operators and uncertainty principle x E Schrodinger wave equation for a particle in one dimensional box Evaluation of average position average momentum and determination of uncertainty in position and momentum and hence Heisenberg's uncertainty principle Pictorial representation of the wave equation of a particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level Lowest energy of the particle Chapter 2 Thermodynamics I Brief resume of first and second Law of thermodynamics Entropy changes in reversible and irreversible processes Variation of entropy with temperature pressure and volume Entropy concept as a measure of unavailable energy and criteria for the spontaneity of reaction Free energy enthalpy functions and their significance criteria for spontaneity of a process Partial molar quantities free energy volume heat concept Gibb's Duhem equation Chapter 3 Chemical Dynamics I Effect of temperature on reaction rates Rate law for opposing reactions of 1st order and IInd order Rate law for consecutive Collision theory of reaction rates and its limitations Steric factor Activated complex theory Ionic reactions single and double sphere models Influence of solvent and ionic strength The comparison of collision and activated complex theory Chapter 4 Electrochemistry I Ion Ion Interactions The Debye Huckel theory of ion ion interactions Potential and excess charge density as a function of distance from the central ion Debye Huckel reciprocal length Ionic cloud and its contribution to the total potential Debye Huckel limiting law of activity coefficients and its limitations Ion size effect on potential Ion size parameter and the theoretical mean activity coefficient in the case of ionic clouds with finite sized ions Debye Huckel Onsager treatment for aqueous solutions and its limitations Debye Huckel Onsager theory for non aqueous solutions The solvent effect on the mobality at infinite dilution Equivalent conductivity vs concentration c 1 2 as a function of the solvent Effect of ion association upon conductivity Debye Huckel Bjerrum equation Chapter 5 Quantum Mechanics II Schrodinger wave equation for a particle in a three dimensional box The concept of degeneracy among energy levels for a particle in three dimensional box Schrodinger wave equation for a linear harmonic oscillator Zero point energy of a particle possessing harmonic motion and its consequence Schrodinger wave equation for three dimensional Rigid rotator Energy of rigid rotator Space

quantization Schrodinger wave equation for hydrogen atom separation of variable in polar spherical coordinates and its solution Principle azimuthal and magnetic quantum numbers and the magnitude of their values Probability distribution function Radial distribution function Shape of atomic orbitals s p Law of mass action and its thermodynamic derivation Third law of thermodynamics Nernest heat theorem determination of absolute entropy unattainability of absolute zero and its limitation Phase diagram for two completely miscible components systems Eutectic systems Calculation of eutectic point Systems forming solid compounds Ax By with congruent and incongruent melting points Phase diagram and thermodynamic treatment of solid solutions Chapter 7 Chemical Dynamics II Chain reactions hydrogen bromine reaction pyrolysis of acetaldehyde decomposition of ethane Photochemical reactions hydrogen bromine General treatment of chain reactions ortho para hydrogen conversion and hydrogen bromine reactions Apparent activation energy of chain reactions Chain length Rice Herzfeld mechanism of organic molecules decomposition acetaldehyde Branching chain reactions and explosions H2 O2 reaction Kinetics of one intermediate enzymatic reaction Michaelis Menton treatment Evaluation of Michaelis s constant for enzyme substrate binding by Lineweaver Burk plot and Eadie Hofstae methods Competitive and non competitive inhibition Chapter 8 Electrochemistry II Ion Transport in Solutions Ionic movement under the influence of an electric field Mobility of ions Ionic drift velocity and its relation with current density Einstein relation between the absolute mobility and diffusion coefficient The Stokes Einstein relation The Nernst Einstein equation Walden's rule The Rate process approach to ionic migration The Rate process equation for equivalent conductivity Total driving force for ionic transport Nernst Planck Flux equation Ionic drift and diffusion potential the Onsager phenomenological equations The basic equation for the diffusion Planck Henderson equation for the diffusion potential Physical Chemistry for the Chemical and Biological Sciences Raymond Chang, 2000-05-12 Hailed by advance reviewers as a kinder gentler P Chem text this book meets the needs of an introductory course on physical chemistry and is an ideal choice for courses geared toward pre medical and life sciences students Physical Chemistry for the Chemical and Biological Sciences offers a wealth of applications to biological problems numerous worked examples and around 1000 chapter end problems Text-Book of Physical Chemistry Clarence Livingston Speyers, 2015-06-26 Excerpt from Text Book of Physical Chemistry Excepting the chapter on Phases this book represents what I have been in the habit of teaching the Senior Class in the Chemical Course at Rutgers College I find lectures very unsatisfactory The subject needs much thought and the student does not seem able to get the proper material for this from notes unless the lectures are made mere dictations Besides Physical Chemistry has attained such development that it is well to put a certain part of it in permanent textbook form now for whatsoever changes in our views time may bring certain ideas we now have will not change These fundamental ideas I have tried to put into this little book Of course I have also included some theories and ideas that later on may have to be rejected That cannot be helped It is difficult to separate that which should be considered in a book of this kind from that which should be excluded I have tried to give the most important

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