

**A. MATTER**

• Mass + Volume

STATES	SOLID	LIQUID	GAS
Shape	Definite *non-compressible	Indefinite *assumes container shape	Indefinite *compressible
Volume		Definite	
Molecular motion	Vibration *2 stones	Gliding *ex. water falls	Constant random

Plasma/ Ionized Gas – 4th state; most abundant state of matter.

- Has p⁺ and e⁻ (thus, greatly affected by magnetic field)
- Ex. ionized Ne light, Aurora, Stars, Sun
- IFA Strength most | e⁻ strongest: S > L > G > P
- Enthalpy (heat/ reaction energy): P > G > L > S

PHASE CHANGE

Melting *aka: Fusion, Liquefaction, Thawing	(Solid to Liquid)
Freezing	(Liquid to Solid)
Evaporation	(Liquid to Gas)
Condensation	(Gas to Liquid)
Sublimation *moth/naphthalene balls	(Solid to Gas)
Deposition *dry ice/ cardice	(Gas to Solid)
Recombination *aka: Deionization	(Plasma to Gas)
Ionization	(Gas to Plasma)

**MATTER CLASSIFICATION****1. Pure substance**

- **Element** – simplest form of substance.
- **Compound** – 2 or more chemical united (separated via chemical means)

2. Mixture – 2 or more substance wherein individual substance identifies are retained (separated via physical means. Alcohol + Water via distillation)

- **Homogeneous** – 1 phase; solution *clear colored
- **Heterogeneous** – 2 phases; suspension, colloid *ex. milk

CLASSIFICATION BASED ON DEPENDENT TO THE AMOUNT OF MATTER PRESENT**1. Extrinsic Property "Dependent"**

- Length, mass/weight, volume, pressure, entropy, enthalpy, electrical resistance

2. Intrinsic Property "Independent"

- Density/ SpGr (water = 1g/ml or cc), viscosity (resistance to flow), velocity (m/sec), temperature, color

B. FUNDAMENTAL CHEMISTRY LAWS**1. Law of Conservation of Mass/ Matter – Antoine Lavoisier**

- Mass/ Matter is always constant (neither created nor destroyed)

2. Law of Definite/ Constant Proportions – Joseph Proust

(Proust's law)

- Chemical compounds always contain the exact proportion of element in fixed ratio (by mass)
- Ex. $\text{H}_2\text{O} \rightarrow 2\text{H} + \text{O}$, $\text{C}_2\text{H}_5\text{O}_2 = \text{CH}_3\text{O}$

3. Law of Multiple Proportion – John Dalton

- When 2 elements form more than 1 compounds, it can be expressed in a fixed whole number (by mass)
- Ex. $\text{CO} \rightarrow 28\text{g/mole}$, $\text{CO}_2 \rightarrow 44\text{g/mole}$
C = 12g/mole
O = 16g/mole

4. Law of combining weights

- Proportions by weight when chemical reaction takes place can be expressed in small integral unit
- Ex. $\text{MgO} \rightarrow 40\text{g/mole}$ (100%)
Mg = 24g/mole (60%); O = 16g/mole (40%)

C. ATOMIC STRUCTURE**1. Democritus – Atoms**

- "Indivisible"

2. John Dalton – Billiard ball

- Matter is made up of atoms

- Postulates:

- Elements are composed of indivisible, indestructible atoms
- Atoms alike for a given element (isotopes)
- Atoms of different elements differ in size, mass & other properties (isobars)
- Compound are formed from 2 or more atoms at different elements
- Atoms combined in simple numerical ratios to form compounds

4. J.J. Thompson – Plum Pudding/ Raisin bread

- e in (+) framework

5. Ernest Rutherford (discoverer of proton) – Nuclear (Gold foil/ a-scattering experiment)

- atom is mostly empty; (+) particles in nucleus

6. Neil Bohr – Planetary

- mostly used

7. Erwin Schrodinger – Quantum/ Mechanical/ e- cloud

- Modern atomic Model; estimates the probability of finding an e in certain position (i.e. at e cloud/ orbital)

Atoms**Proton = (+) ion**

- Atomic number (basis of electronic configuration)
- **Ernest Rutherford**

Electrons = (-) ion

- p⁺ in uncharged state
- negligible weight 1.836x lighter than p⁺
- **J.J. Thompson**
- Cathode ray tube: e⁻ m/2 ratio
- **R.A. Millikan**
- Oil drop experiment: measure accurate charge and mass of e⁻

Student Models And Exercises For General Chemistry

C Cleary



Student Models And Exercises For General Chemistry:

Modeling Chemical Systems using Cellular Automata Lemont B. Kier, Paul G. Seybold, Chao-Kun Cheng, 2006-02-23

Modeling Chemical Systems using Cellular Automata provides a practical introduction to an exciting modeling paradigm for complex systems. The book first discusses the nature of scientific inquiry using models and simulations and then describes the nature of cellular automata models. It then gives detailed descriptions with examples and exercises of how cellular automata models can be used in the study of a wide variety of chemical, physical, and biochemical phenomena. Topics covered include models of water itself, solution phenomena, solution interactions with stationary systems, first and second order kinetic phenomena, enzyme kinetics, vapor liquid equilibrium, and atomic and molecular excited state kinetics. The student experiences these systems through hands-on examples and guided studies. This book is the first of its kind, a textbook and a laboratory manual about cellular automata modeling of common systems in chemistry. The book is designed to be used as a text in undergraduate courses dealing with complex systems and/or as a computational supplement to laboratory courses taught at the undergraduate level. The book includes compact descriptions of a large variety of physical and chemical phenomena, illustrative examples of simulations with exercises for further study, an instructor's manual for use of the program. The book will be of great value in undergraduate courses in chemistry, physics, biology, applied mathematics, and bioinformatics and as a supplement for laboratory courses in introductory chemistry, organic chemistry, physical chemistry, medicinal chemistry, chemical engineering, and other courses dealing with statistical and dynamic systems. It allows the exploration of a wide range of dynamic phenomena, many of which are not normally accessible within conventional laboratory settings due to limitations of time, cost, and experimental equipment. The book is both a textbook on applied Cellular Automata and a lab manual for chemistry, physics, engineering courses with lab activity. It would supplement other lab work and be an additional book the students would use in the course. The authors have assessed the emerging need for this kind of activity in science labs because of the cost of the practical activities and the frequent failure of some exercises leading to lost didactic value of some experiments. This book is pioneering an alternative that will grow in use. There are no course directors who would use Cellular Automata exclusively. The authors see an emerging interest in this kind of work in courses that contain lab exercises. One such course is the graduate course that Lemont Kier gives in Life Sciences about complexity. He uses many examples and studies from Cellular Automata in the latter part of this course.

Problems and Problem Solving in Chemistry Education Georgios Tsapralis, 2021-05-17

Problem solving is central to the teaching and learning of chemistry at secondary, tertiary, and post tertiary levels of education, opening to students and professional chemists alike a whole new world for analysing data, looking for patterns, and making deductions. As an important higher order thinking skill, problem solving also constitutes a major research field in science education. Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem

solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner *Catalogue of the Officers and Students of Columbia College, for the Year ...* Columbia College (New York, N.Y.),1869 *Simulations and Student Learning* Matthew Schnurr,Anna MacLeod,2021-01-04 The book underlines the value of simulation based education as an approach that fosters authentic engagement and deep learning

Making Chemistry Relevant Sharmistha Basu-Dutt,2010-02-19 Unique new approaches for making chemistry accessible to diverse students Students interest and achievement in academics improve dramatically when they make connections between what they are learning and the potential uses of that knowledge i n the workplace and or in the world at large Making Chemistry Relevant presents a unique collection of strategies that have been used successfully in chemistry classrooms to create a learner sensitive environment that enhances academic achievement and social competence of students Rejecting rote memorization the book proposes a cognitive constructivist philosophy that casts the teacher as a facilitator helping students to construct solutions to problems Written by chemistry professors and research groups from a wide variety of colleges and universities the book offers a number of creative ways to make chemistry relevant to the student including Teaching science in the context of major life issues and STEM professions Relating chemistry to current events such as global warming pollution and terrorism Integrating science research into the undergraduate laboratory curriculum Enriching the learning experience for students with a variety of learning styles as well as accommodating the visually challenged students Using media hypermedia games and puzzles in the teaching of chemistry Both novice and experienced faculty alike will find valuable ideas ready to be applied and adapted to enhance the learning experience of all their students *Senate documents* ,1883 *Industrial Education in the United States* United States. Bureau of Education,1883 *Catalogue* George Washington University,1907 **Catalogue of the Columbian College in the District of Columbia** Columbian College in the District of Columbia,1915 **Senate Documents, Otherwise Publ. as Public Documents and Executive Documents** United States. Congress. Senate,1883

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