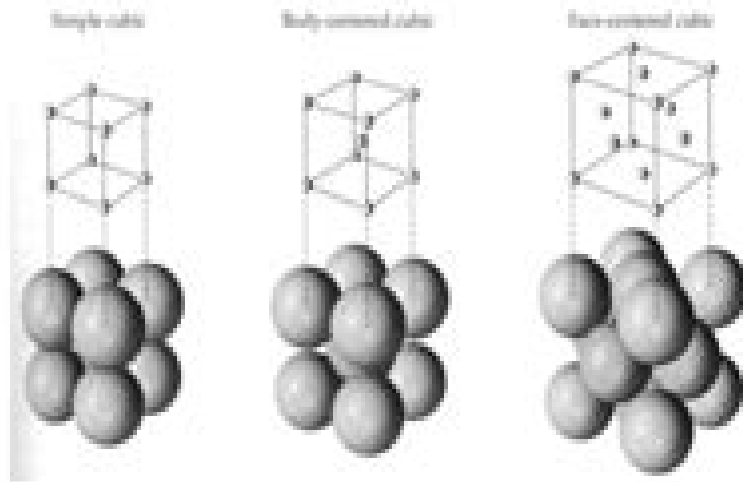


In the given figure below, simple cube is a primitive cell. Number of atoms per unit cell is one for it. Body centered cubic and face centered cubic is non primitive. No. of atoms per unit cell is 2 and 4 respectively.



Simple Cubic Structure (SC)

Simple cube has eight corner atoms. Each corner atom contribute $1/8^{\text{th}}$ to unit cell. And each atom has six nearest atom.

Number of atoms per unit cell = 8 atoms \times $1/8 = 1$

Co-ordination number = 6



If a is lattice constant then atomic radius, $r = a/2$

Atomic packing fraction = $\frac{\text{volume of atoms}}{\text{volume of unit cell}}$

$$\begin{aligned} \text{Atomic packing fraction} &= \frac{1 \times \frac{4}{3}\pi r^3}{a^3} = \frac{\frac{4}{3}\pi r^3}{(2r)^3} = \frac{\frac{4}{3}\pi r^3}{8r^3} \\ &= \frac{\pi}{6} = 0.526 = 52.6\% \end{aligned}$$

Body Centered Cubic Structure (BCC)

Body centered cube has eight corner atoms and one atom at the center. Each corner atom contribute $1/8^{\text{th}}$ to unit cell. And center atom has eight nearest atom.

Number of atoms per unit cell = 8 atoms \times $1/8$ (corner atom) + 1 (center atom) = 2

Co-ordination number = 8

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Table of Contents Solid State Physics Volume 6 Methods Of Expe

1. Understanding the eBook Solid State Physics Volume 6 Methods Of Expe
 - The Rise of Digital Reading Solid State Physics Volume 6 Methods Of Expe
 - Advantages of eBooks Over Traditional Books
2. Identifying Solid State Physics Volume 6 Methods Of Expe
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solid State Physics Volume 6 Methods Of Expe
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solid State Physics Volume 6 Methods Of Expe
 - Personalized Recommendations
 - Solid State Physics Volume 6 Methods Of Expe User Reviews and Ratings
 - Solid State Physics Volume 6 Methods Of Expe and Bestseller Lists
5. Accessing Solid State Physics Volume 6 Methods Of Expe Free and Paid eBooks
 - Solid State Physics Volume 6 Methods Of Expe Public Domain eBooks
 - Solid State Physics Volume 6 Methods Of Expe eBook Subscription Services
 - Solid State Physics Volume 6 Methods Of Expe Budget-Friendly Options
6. Navigating Solid State Physics Volume 6 Methods Of Expe eBook Formats
 - ePub, PDF, MOBI, and More
 - Solid State Physics Volume 6 Methods Of Expe Compatibility with Devices
 - Solid State Physics Volume 6 Methods Of Expe Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solid State Physics Volume 6 Methods Of Expe
 - Highlighting and Note-Taking Solid State Physics Volume 6 Methods Of Expe
 - Interactive Elements Solid State Physics Volume 6 Methods Of Expe
8. Staying Engaged with Solid State Physics Volume 6 Methods Of Expe

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solid State Physics Volume 6 Methods Of Expe
9. Balancing eBooks and Physical Books Solid State Physics Volume 6 Methods Of Expe
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solid State Physics Volume 6 Methods Of Expe
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solid State Physics Volume 6 Methods Of Expe
- Setting Reading Goals Solid State Physics Volume 6 Methods Of Expe
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solid State Physics Volume 6 Methods Of Expe
- Fact-Checking eBook Content of Solid State Physics Volume 6 Methods Of Expe
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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