

SKETCHES OF AUSTRALIAN LIFE & SCENERY. COMPLETE IN 12 PLATES.



# Sketches Of Australian Life Scenery

**R Barnett**



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web 10 flight attendants and 203 passengers on board all 217 people on board were killed and the airplane was destroyed

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{O}_2(\text{g}) + \text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$  from the following data  
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ/mole}$

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web hess s law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the  $\Delta h$  values for each step to use hess s law two principles must be understood one if an equation is reversed the sign of the  $\Delta h$  value is also reversed two if an equation is multiplied by a

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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multiply the mass by the change in temperature and the  $c$  given trial 1 use Hess's law and the following equations and  $\Delta H$  values to determine the heat of reaction for the reaction  $c_2H_4$

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web thus for reaction 9.2 in which 1 mole of  $H_2O$  and 1 mole of  $OH^-$  combine to form 2 moles of  $H_2O$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove Hess's law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 Hess's law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $1. NaOH(s) \rightarrow Na^+(aq) + OH^-(aq)$   $2. NaOH(s) \rightarrow H^+(aq) + Cl^-(aq)$   $3. H_2O(l) \rightarrow H^+(aq) + Cl^-(aq)$