



THE FABRIC & NEEDLECRAFT PROJECT BOOK



*70 easy-to-make projects and over
350 step-by-step photographs*

Contributing Editor: Penelope Cream



The Fabric Needlecraft Project

Andria Lisle, Amie Petronis Plumley



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Sewing School ® Andria Lisle,Amie Petronis Plumley,2015-08-28 Kids can sew too This inspiring guide includes 21 fun sewing projects for children ages 5 and up With easy to follow illustrated instructions and cut out patterns young crafters will quickly be sewing up colorful pillows potholders dolls blankets and more These kid tested projects require only minimal supervision and most can be made using simple hand stitches so no sewing machine is required With plenty of encouragement and helpful tips Sewing School helps young sewers develop their skills while sparking a passion that will last a lifetime Also available in this series Sewing School 2 Sewing School Fashion Design Sewing School Quilts and Sewing School Box Set

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 what is the limiting reagent if 8.52 g are actually formed what is the

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 step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.6656 \text{ moles O}_2$

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 yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction Al + O₂ → Al₂O₃
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.926 \text{ mol Al}$
 $\frac{5.926 \text{ mol Al}}{4} \times 2 = 2.963 \text{ mol Al}_2\text{O}_3$
 $2.963 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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nominal 0 2 g 10 min up to 1 4 g 10 min din en iso 1872 1 condition t which meet the requirements according to din 8075 this
standard is also valid for the dimensions and operating pressures of pipes made of coextruded layers