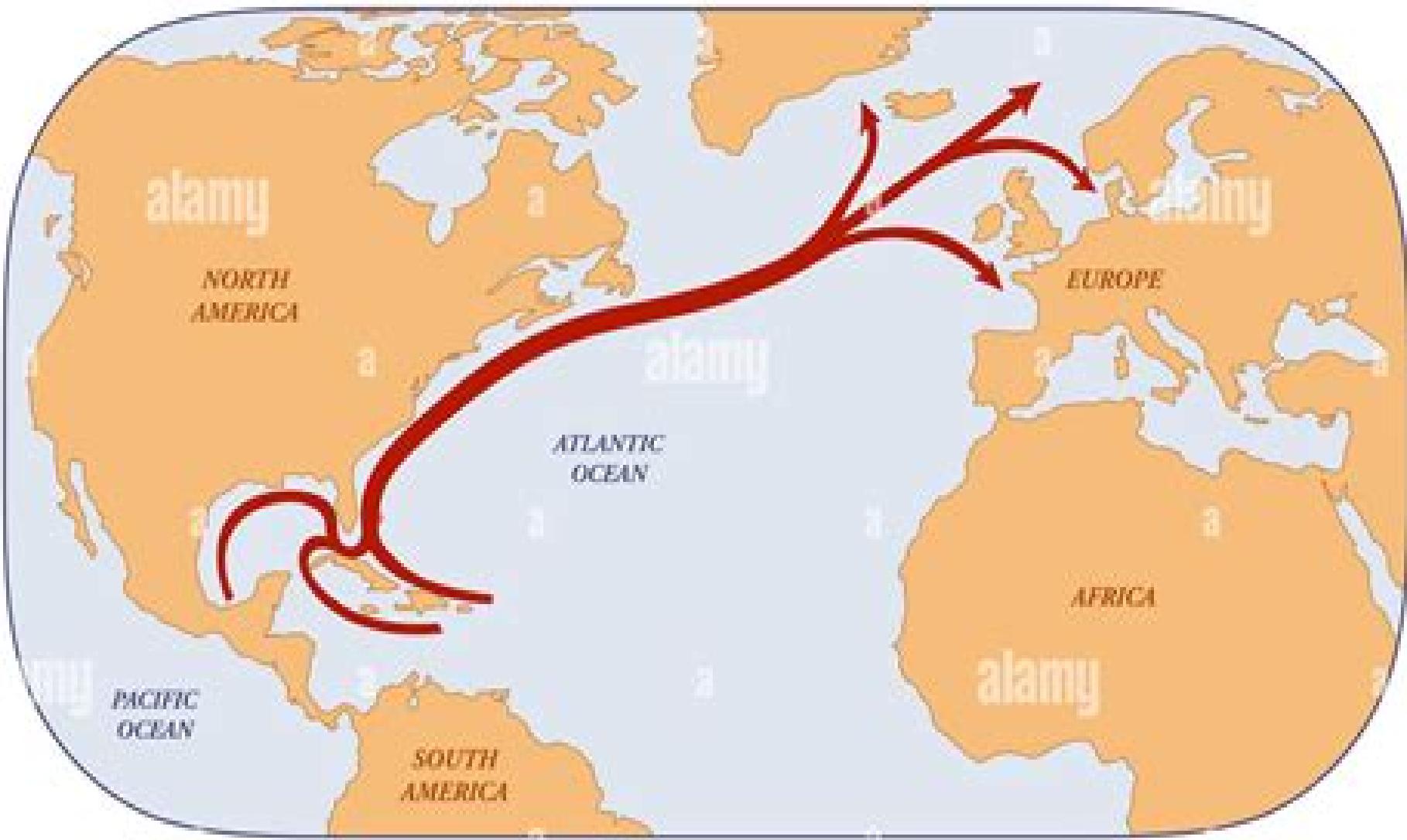


# The Gulf Stream



# The Gulf Stream

**August Heinrich Petermann, W. von  
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0.00581158 molal  $= \frac{0.00581158 \text{ mol}}{0.4500 \text{ kg}} = 0.0129 \text{ mol/l}$  mole fraction  $n_A = \frac{0.825}{0.825 + 0.00581158} = 0.825$   
 $n_B = \frac{0.00581158}{0.825 + 0.00581158} = 0.00581158$  molality  $0.825 \text{ g} / 1.0 \text{ kg} = 0.825 \text{ molal}$   
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 $n_B = \frac{0.4500}{0.00581158 + 0.4500} = 0.4500$  g/mol  $18.015 \text{ g/mol} = 24.97918401 \text{ molal}$

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calculate molarity 5 moles / 0.25 liters = 20 M 3 c a solution with molarity 2 requires 2 mol of NaOH per liter so  $2 \text{ mol} / 0.25 \text{ L} = 8 \text{ mol/L}$   
a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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 $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2}{V_1 + V_2}$  where  $M_1$  and  $M_2$  are the molarities of the two components and  $V_1$  and  $V_2$  are their respective volumes. If the total volume of the mixture is  $V$ , then the total molarity is  $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2}{V}$ .  
with increase in temperature which of these changes a molality b fraction of solute present in water

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