

Particle statistics

In a famous series of calculations published in Klug and Alexander it was calculated that IF a powder was crushed to 1 micron in size, the amount of powder in a large cavity mount, would be barely sufficient to achieve 1% reproducibility in intensity.

This calculation assumed a single point detector having a small detection volume which in this example, at 1 micron, would detect 38,000 out of 38 billion particles in the sampling volume. Notice the dramatic increase in available particles as the size is decreased.

How many particles do you need for 1% accuracy in intensity

	<u>40μm</u>	<u>10μm</u>	<u>1μm</u>
# diffraction particles	12	760	38000

To achieve 1% accuracy

$$\sigma = \sqrt{n}/n$$

$$\text{Std. err.} = 2.3 \sigma < 1\%$$

$$n > 52,900 \leftarrow$$



Small Particle Statistics An Account Of

E.C. Monahan, Gearóid Mac Niocaill



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jan 3 2023 a net ionic equation is a chemical equation that shows only the species participating in a chemical reaction and omits the spectator ions spectator ions are ions that

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4^+ \text{aq} + 2\text{F}^- \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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