

THE  
POOR-MAN'S  
COMFORT.

A  
Tragi-Comedy,

As it was divers times Acted  
in the Theatre in Drury Lane  
with great applause.

Written by  
ROBERT DAINBORNE  
Master of Arts.



LONDON:  
Printed for John Johnson at the Golden Ball  
in St. Dunstons Church-yard, and John Baskin at  
the Angel in Pall-mall Alley.  
1733.

# The Poor Mans Comfort

**S Ashworth**



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 $\text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$   
 $2 \text{NaF} + \text{Cl}_2 \rightarrow 2 \text{H}_2 + \text{O}_2$   
 $\text{H}_2\text{O} + \text{Pb} + \text{OH}^- \rightarrow 2 \text{HCl} + \text{H}_2\text{O} + \text{PbCl}_2$   
 $\text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + \text{Al}_2\text{SO}_4 + 3 \text{CH}_4 + 2 \text{O}_2$   
 $\text{CO}_2 + 2 \text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 5 \text{O}_2$   
 $3 \text{CO}_2 + 4 \text{H}_2\text{O} \rightarrow \text{C}_8\text{H}_{18} + 25 \text{O}_2$   
 $16 \text{CO}_2 + 18 \text{H}_2\text{O}$

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 solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas  
 gaseous butane  $\text{C}_4\text{H}_{10}$  reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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 hen balancing  
 $1 \text{Na} + 3 \text{PO}_4 + 3 \text{KOH} + 3 \text{NaOH} + 1 \text{K} + 3 \text{PO}_4 + 1 \text{MgF}_2 + 1 \text{Li} + 2 \text{CO}_3 + 1 \text{MgCO}_3 + 2 \text{LiF} + 1 \text{P}_4 + 3 \text{O}_2 + 2 \text{P}_2\text{O}_3 + 2 \text{RbNO}_3 + 1 \text{BeF}_2 + 1 \text{BeNO}_3 + 2 \text{RbF}$

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 $1 \text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$   
 $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$   
 $2 \text{NaF} + \text{Cl}_2 \rightarrow 2 \text{H}_2 + \text{O}_2$   
 $5 \text{H}_2\text{O} + \text{Pb} + \text{OH}^- \rightarrow 2 \text{HCl} + 2 \text{H}_2\text{O} + \text{PbCl}_2$   
 $2 \text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + \text{Al}_2\text{SO}_4 + 3 \text{CH}_4 + 2 \text{O}_2$   
 $1 \text{CO}_2 + 2 \text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 5 \text{O}_2$   
 $3 \text{CO}_2 + 4 \text{H}_2\text{O}$

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 $1 \text{O}_2 + 2 \text{NH}_3 \rightarrow \text{HNO}_3 + \text{H}_2\text{O} + 2 \text{O}_2 + 3 \text{H}_2$

2 o 2 h 2 o o 2 4 rb h 2 o rboh h 2 5 cl 2 kbr kcl br 2 6 n 2 h 2 nh 3 7 ndf 3 ca nd caf 2 8 ch 4 h 2 o co h 2 9 c 2 h 5 oh

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products and direction of the reaction balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow

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