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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $\text{Mg}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$
 $2\text{NaClO}_3 \rightarrow 2\text{NaCl} + 3\text{O}_2$
 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 $\text{N}_2 + \text{O}_2 \rightarrow 2\text{NO}$
 $2\text{NH}_3 + 3\text{O}_2 \rightarrow 2\text{N}_2 + 6\text{H}_2\text{O}$

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 $\text{Mg}(\text{OH})_2 + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$
 note all reactants and products require a coefficient of at least one
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 metal acid salt hydrogen

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 $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$
 if 1.0 mol of C_2H_2 is burned how many moles of CO_2 are formed
 2.5 mol C_2H_2 is burned how many moles of oxygen gas are consumed
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