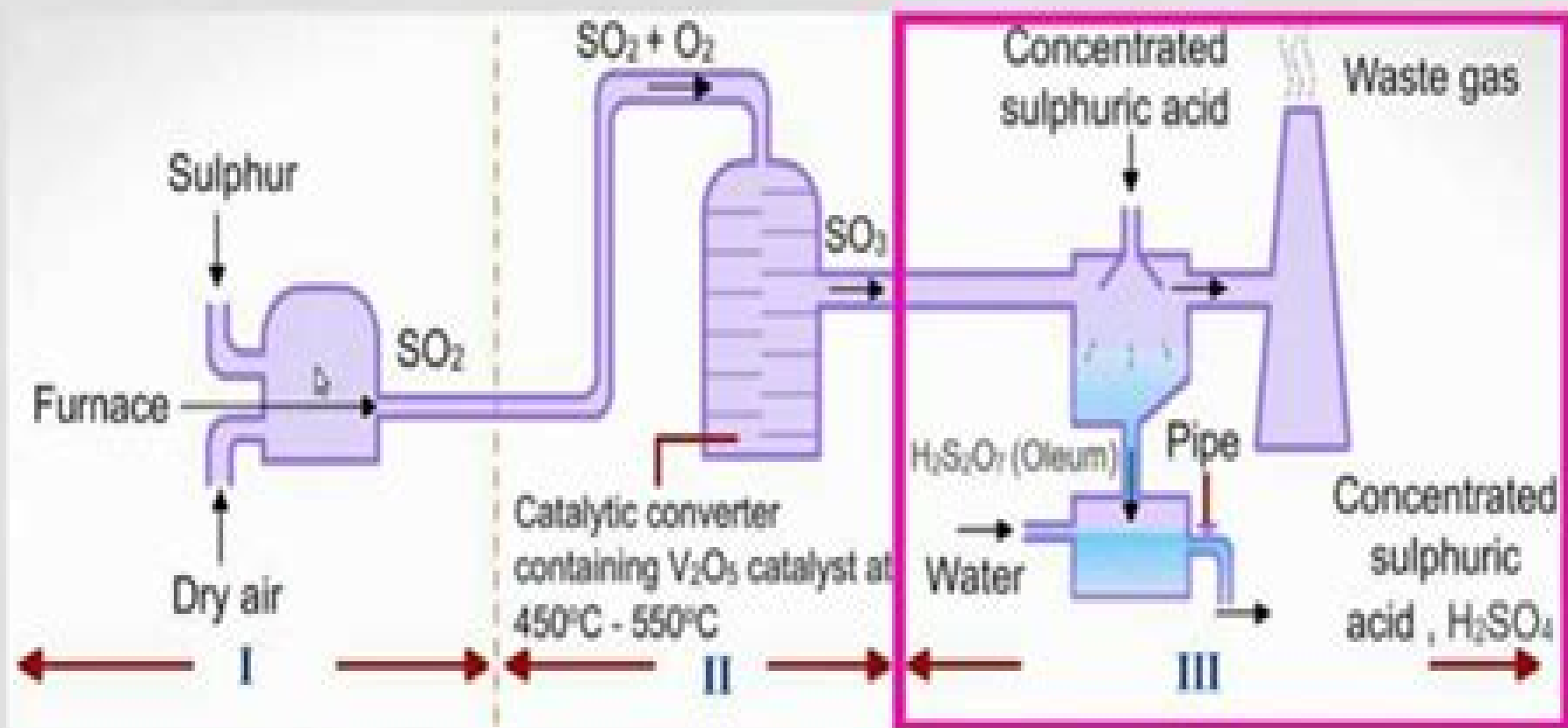


Manufacture of sulphuric acid in industry

►► Contact process



Sulphuric Acid Manufacture

**Matt King, Michael Moats, William G.
Davenport, Matthew J. King**



Sulphuric Acid Manufacture:

Handbook of Sulphuric Acid Manufacturing Douglas K. Louie, 2005 *Sulfuric Acid Manufacture* Matt King, Michael Moats, William G. Davenport, Matthew J. King, 2005-11-04 More sulfuric acid is produced every year than any other chemical. It has a wide range of uses including phosphate fertilizer production, explosives, glue, wood preservatives, and lead acid batteries. It is also a particularly corrosive and dangerous acid with extreme environmental and health hazards if not manufactured, used, and regulated properly. *Sulfuric Acid Manufacture: Analysis, Control, and Optimization* keeps the important topics of safety and regulation at the forefront as it overviews and analyzes the process of sulfuric acid manufacture. The first nine chapters focus on the chemical plant processes involved in industrial acidmaking with considerable data input from the authors' industrial colleagues. The last 15 chapters are dedicated to the mathematical analysis of acidmaking. Both authors bring years of hands-on knowledge and experience to the work, making it an exceptional reference for anyone involved in sulfuric acid research and/or manufacture. Only book to examine the processes of sulfuric acid manufacture from an industrial plant standpoint as well as mathematical. Draws on the industrial connections of the authors through their years of hands-on experience in sulfuric acid manufacture. A considerable amount of industrial plant data is presented to support the text.

Sulfuric Acid Manufacture Matt King, Michael Moats, William G. Davenport, 2013-05-11 By some measure the most widely produced chemical in the world today, sulfuric acid has an extraordinary range of modern uses including phosphate fertilizer production, explosives, glue, wood preservative, and lead acid batteries. An exceptionally corrosive and dangerous acid, production of sulfuric acid requires stringent adherence to environmental regulatory guidance within cost-efficient standards of production. This work provides an experience-based review of how sulfuric acid plants work, how they should be designed, and how they should be operated for maximum sulfur capture and minimum environmental impact. Using a combination of practical experience and deep physical analysis, Davenport and King review sulfur manufacturing in the contemporary world where regulatory guidance is becoming ever tighter and where new processes are being required to meet them and where water consumption and energy considerations are being brought to bear on sulfuric acid plant operations. This 2e will examine in particular newly developed acid-making processes and new methods of minimizing unwanted sulfur emissions. The target readers are recently graduated science and engineering students who are entering the chemical industry and experienced professionals within chemical plant design companies, chemical plant production companies, sulfuric acid recycling companies, and sulfuric acid users. They will use the book to design, control, optimize, and operate sulfuric acid plants around the world. Unique mathematical analysis of sulfuric acid manufacturing processes providing a sound basis for optimizing sulfuric acid manufacturing processes. Analysis of recently developed sulfuric acid manufacturing techniques suggests advantages and disadvantages of the new processes from the energy and environmental points of view. Analysis of tail gas sulfur capture processes indicates the best way to combine sulfuric acid making and tailgas sulfur capture processes.

from the energy and environmental points of view Draws on industrial connections of the authors through years of hands on experience in sulfuric acid manufacture *The Manufacture of Sulphuric Acid in the United States* Arthur Edward Wells,Donald E. Fogg,1920 Sulfuric Acid Manufacture Andrew Miller Fairlie,1936 **A Practical Guide to the Manufacture of Sulfuric Acid, Oleums, and Sulfonating Agents** Navin G. Ashar,Kiran R. Golwalkar,2013-11-23 This critical volume provides practical insights on sulfuric acid and related plant design and on techniques to improve and enhance substantially the efficiency of an existing plant by means of small modifications The book provides readers with a better understanding of the state of art in sulfuric acid manufacture as well as importantly in the manufacture of value added products based on sulfur that are also associated with the manufacture of sulfuric acid Overall engineers and plant managers will be introduced to technologies for making their sulfuric acid enterprises more productive remunerative and environmentally friendly A Practical Guide to the Manufacture of Sulfuric Acid Oleums and Sulfonating Agents covers sulfuric acid and derivative chemical plant details from the nuts and bolts level to a holistic perspective based on actual field experience The book is indispensable to anyone involved in implementing a sulfuric acid or related chemical plant *The Manufacture of Sulfuric Acid* Werner W Duecker (ed.),James Robert West,1959 The Chemistry of Sulphuric Acid Manufacture Henry Arthur Smith,1873 The Manufacture of Sulphuric Acid in the United States Arthur Edward Wells,Donald E. Fogg,1920 **The Manufacture of Sulphuric Acid (contact Process)** Frank Douglas Miles,1925

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton $h_a + b \rightleftharpoons h_b$ strong acids e g sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web jul 19 2023 at the beginning of the titration the solution in the erlenmeyer flask is acidic as the base is added it completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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