

Solute-Solvent Interactions

- Solubility varies greatly with the type of compounds involved.
- “Like dissolves like” is a rough but useful rule for predicting whether one substance will dissolve in another.
- What makes substances similar depends on:
 - type of bonding
 - polarity or nonpolarity of molecules
 - intermolecular forces between the solute and solvent



Solute Solvent Interactions

**Johannes François Coetzee, Calvin D.
Ritchie**



Solute Solvent Interactions:

Solute-solvent Interactions Johannes Francois Coetzee, Calvin D. Ritchie, 1969 Highlights in Solute-Solvent Interactions Wolfgang Linert, 2012-12-06 Most organic molecules retain their integrity when dissolved and even though in such cases the effects exerted by solvents are in the language of the coordination chemist of the outer sphere kind the choice of solvent can be critical to the successful outcome of an operation or preparation Solubilities of reactants and products must be taken into account and even if the organic principals in the reactions retain their integrity many of the reagents are electrolytes and their state of aggregation will affect their reactivity In testifying to the importance of understanding solute solvent interactions I draw attention to a large class of inorganic species for which the involvement in the chemical and physical properties by the solvent is even more deeply seated It is comprised by the large body of metal atoms in low oxidation states for which solvent molecules intervene as reagents At the same time because the ions carry charges the effects arising from outer sphere interactions are usually greater than they are for neutral molecules To cite an example when FeCl_3 is dissolved in water to form a dilute say 0.01 M solution there is a complete reorganization of the coordination sphere of the cation Whereas in the solid each cation is surrounded by six chloride ions in the solution the dominant form is $\text{Fe}(\text{H}_2\text{O})_6^{3+}$ followed by $\text{Fe}(\text{H}_2\text{O})_5\text{Cl}^{2+}$ $\text{Fe}(\text{H}_2\text{O})_4\text{Cl}_2$ etc in rapidly decreasing abundance Solute-Solvent Interactions, 1969 Johannes François Coetzee, Calvin D. Ritchie, 1969 Quantitative Treatments of Solute/Solvent Interactions P. Politzer, J.S. Murray, 1994-12-07 The primary objective of this volume the first in a new series entitled Theoretical and Computational Chemistry is to survey some effective approaches to understanding describing and predicting ways in which solutes and solvents interact and the effects they have upon each other The treatment of solute solvent interactions that is presented emphasizes a synergism between theory and experiment Data obtained experimentally are used as a basis for developing quantitative theoretical models that permit the correlation and interpretation of the data and also provide a predictive capability The latter being of course a key motivation for these efforts Linear solvation energy relationships have been quite successful in this respect and accordingly receive considerable attention Other effective approaches including computational ones are also being pursued and are discussed in several chapters This is an area that is continually evolving and it is hoped that the present volume will convey a sense of its dynamic nature *Solute-solvent interactions*, ed Johannes F. Coetzee, Calvin D. Ritchie, **Solute solvent interactions in chemistry and biology** Mortimer J. Kamlet, Michael H. Abraham, 1989 **Solute-solvent Interactions** Johannes Francois Coetzee, Calvin D. Ritchie, 1969 **A general model for the treatment of solute-solvent interactions** Jose Luis Escobar Valderrama, 1996 Solute Solute Solvent Interactions, 1981 **Plenary and Invited Lectures Presented at the Eighth International Symposium on Solute Solute Solvent Interactions** International Symposium on Solute Solute Solvent Interactions. 8, 1987, Regensburg, 1987 **Protein Purification Process Engineering** Roger Harrison, 2019-07-16 Offers coverage of the development of protein

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