

PROBLEMS/SOLUTIONS

1. The mixing ratio of oxygen in the atmosphere is 20.95%. Calculate the concentration in mol L⁻¹ and in g m⁻³ at P° (101 325 Pa, 1.00 atm) and 25°C.

Solution

The mixing ratio for O₂ (g) is 20.95% (Table 2.1). Use the 'Ideal Gas' law (PV = nRT) to calculate the total number of moles of gas in 1.00 L for the given conditions. Constants are given in Appendices B.1 and C.1.

T = 25°C (298 K), P = 1.01325 x 10⁵ Pa, R = 8.315 J K⁻¹ mol⁻¹, and V = 1.00 L = 1.00 x 10⁻³ m³

The total number of moles of gas in 1.00 L is

$$\begin{aligned} n &= PV/RT \\ n &= 0.04089 \text{ mol} \end{aligned}$$

The number of moles of oxygen can be calculated from the mole fraction (which is the same as the mixing ratio).

$$\% \text{ O}_2 = 20.95, \text{ mole fraction} = 0.2095$$

The number of moles of O₂ in 1.00 L is: 0.2095 x 0.04089 = 0.00857 mol

The concentration of O₂ (g) is: 8.57 x 10⁻³ mol L⁻¹.

The concentration of O₂ (g) in units of g m⁻³ is determined as follows:

1 m³ = 1000 L, therefore, 1000 L contains 8.57 moles of O₂ (g). (M.M. of O₂ is 31.9988 g mol⁻¹)

$$31.9988 \text{ g mol}^{-1} \times 8.57 \text{ mol m}^{-3} = 274 \text{ g m}^{-3}$$

The O₂ (g) concentration can be expressed as either 8.57 x 10⁻³ mol L⁻¹ or 274 g m⁻³.

2. Calculate the atmospheric pressure at the stratopause. What are the concentrations (mol m⁻³) of dioxygen and dinitrogen at this altitude? How do these concentrations compare with the corresponding values at sea level?

Solution

Use Equation 2.3 to calculate the pressure at 50 km (stratopause).

$$P_h = P^\circ e^{-\bar{M}_a g h}$$

R = 8.315 J K⁻¹ mol⁻¹, P° = 101 325 Pa, g = 9.81 m s⁻², $\bar{M}_a$ = 0.02896 kg mol⁻¹ (average molar mass of air), h = 50 km (50 000 m) and T = -2°C (271 K). Note: these last two parameters are the approximate altitude and temperature at the stratopause.

Solutions Manual For Environmental Chemistry

Colin Baird, Michael Cann



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Table of Contents Solutions Manual For Environmental Chemistry

1. Understanding the eBook Solutions Manual For Environmental Chemistry
 - The Rise of Digital Reading Solutions Manual For Environmental Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Manual For Environmental Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Manual For Environmental Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Manual For Environmental Chemistry

- Personalized Recommendations
- Solutions Manual For Environmental Chemistry User Reviews and Ratings
- Solutions Manual For Environmental Chemistry and Bestseller Lists
- 5. Accessing Solutions Manual For Environmental Chemistry Free and Paid eBooks
 - Solutions Manual For Environmental Chemistry Public Domain eBooks
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 - Solutions Manual For Environmental Chemistry Budget-Friendly Options
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 - Solutions Manual For Environmental Chemistry Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Solutions Manual For Environmental Chemistry
 - Highlighting and Note-Taking Solutions Manual For Environmental Chemistry
 - Interactive Elements Solutions Manual For Environmental Chemistry
- 8. Staying Engaged with Solutions Manual For Environmental Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Manual For Environmental Chemistry
- 9. Balancing eBooks and Physical Books Solutions Manual For Environmental Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Manual For Environmental Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solutions Manual For Environmental Chemistry
 - Setting Reading Goals Solutions Manual For Environmental Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solutions Manual For Environmental Chemistry

- Fact-Checking eBook Content of Solutions Manual For Environmental Chemistry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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