

Summer Of `98: When Homers Flew, Records Fell, And Baseball Reclaimed America

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Summer Of 98 When Homers Flew Records Fell And Baseball Reclaimed America

Marcel A. Müller



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation
$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

8.3 mole $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield
$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
$$\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mole}} = 0.666 \text{ moles } \text{O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained
$$\frac{\text{calculated mass of the product}}{\text{mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al_2O_3
$$\frac{160.0 \text{ g } \text{Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol } \text{Al}}{2 \text{ mol } \text{Al}_2\text{O}_3} \times \frac{101.96 \text{ g } \text{Al}_2\text{O}_3}{1 \text{ mol } \text{Al}_2\text{O}_3} = 297.6 \text{ g } \text{Al}_2\text{O}_3$$

c determine the percent yield

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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