Henderson Hasselbalch Equation

Dissociation of an Acid

$$HA + H_2O \rightleftharpoons A^{\ominus} + H_3O^{\oplus}$$
Acid Water Conjugate Hydronium
Base

Equation

$$pH = pK_a + log_{10} \frac{[A^{\ominus}]}{[HA]}$$

pH:pH value of the solution

K : Dissociation constant of the acid

 $[A^{\Theta}]$: Concentration of the conjugate base (A^{Θ})

[HA]: Concentration of the acid (HA)

The Henderson Equation

Mr. Rohit Manglik

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