

THE CHALLENGE OF JAPAN'S INTERNATIONALIZATION:

ORGANIZATION AND CULTURE

edited by Hiroshi Mannari and Harumi Betsu

The Challenge Of Japans Internationalization Organization And Culture

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gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{Ag} + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{h}^+ \text{aq} + \text{oh}^- \text{aq} \rightarrow \text{h}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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 $2\text{Cl}^{-}(\text{aq}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{CuCl}_2(\text{s})$
 $2\text{Cl}^{-}(\text{aq}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{CuCl}_2(\text{s})$ these spectator ions are not actively involved in the reaction the existence of
spectator ions affects the debye length