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**Saikō Ema**

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the  $\text{OH}^-$  basic solution one in which the  $\text{H}^+$  is less than the  $\text{OH}^-$  alkaline solutions the  $\text{H}^+$  of a basic solution is less than  $1 \times 10^{-7} \text{ M}$

they are the same as basic solutions

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thus ruling out that  $\text{S}^{2-}$  is the weakest base

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electrons during a reaction while a base donates a pair of electrons

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