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rightarrow ag 2cr 2o 7 s label 4 2 3

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web jan 29 2020 the ionic form of this reaction is 2 na aq 2 cl aq cu 2 aq so 4²⁻ aq 2 na aq so 4²⁻ aq cucl 2 s the sodium ions

and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the

equation

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 net ionic equation $\text{h}^+ \text{aq} + \text{cl}^- \text{aq} + \text{na}^+ \text{aq} + \text{hco}_3^- \text{aq} + \text{na}^+ \text{aq} + \text{cl}^- \text{aq} + \text{h}_2\text{o} \text{l} + \text{co}_2 \text{g} \rightarrow \text{h}^+ \text{aq} + \text{hco}_3^- \text{aq} + \text{h}_2\text{o} \text{l} + \text{co}_2 \text{g}$ here is a video on spectator ions
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web the six no_3^- ions and the six na^+ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{ba}^{2+} \text{aq} + 2\text{po}_4^{3-} \text{aq} \rightarrow \text{ba}_3\text{po}_4\text{s}$

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{h}^+ \text{aq} + \text{oh}^- \text{aq} \rightarrow \text{h}_2\text{o} \text{l}$ the spectator ions in $\text{koh} \text{aq}$ and $\text{hno}_3 \text{aq}$ are $\text{k}^+ \text{aq}$ and $\text{no}_3^- \text{aq}$ how to identify and list the

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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