



Site Carpentry And Joinery

David Kirk



Site Carpentry And Joinery:

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web b 3 torching methods follow manufacturer s recommendations on torching to insulation and recover board surface use a small torch when flashing near details do not lay an operating torch directly onto the membranes b there is a danger both of fire and of damaging the membrane

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subsequent Δh values $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{O}_2 \text{ g}$ $\Delta h = 94.5 \text{ kJ}$ $\text{H}_2\text{O g} + \text{H}_2 \text{ g} \rightarrow \frac{1}{2} \text{O}_2 \text{ g}$ $\Delta h = 71.2 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + 3 \text{H}_2\text{O g}$ $\Delta h = 283 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g}$

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web chemistry 120 hess s law worksheet calculate Δh for the reaction $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ from the following data $\text{C}_2\text{H}_4 \text{ g} + 3 \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 2 \text{H}_2\text{O l}$ $\Delta h = 1411 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 3 \text{H}_2\text{O l}$ $\Delta h = 1560 \text{ kJ}$ $\text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O l}$ $\Delta h = 285 \text{ kJ}$ mole

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} + \text{Na}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$ 2 $\text{NaOH (s)} + \text{H}^+ \text{aq} \rightarrow \text{H}_2\text{O (l)}$ 3 $\text{NaOH (s)} + \text{H}^+ \text{aq} \rightarrow \text{H}_2\text{O (l)}$