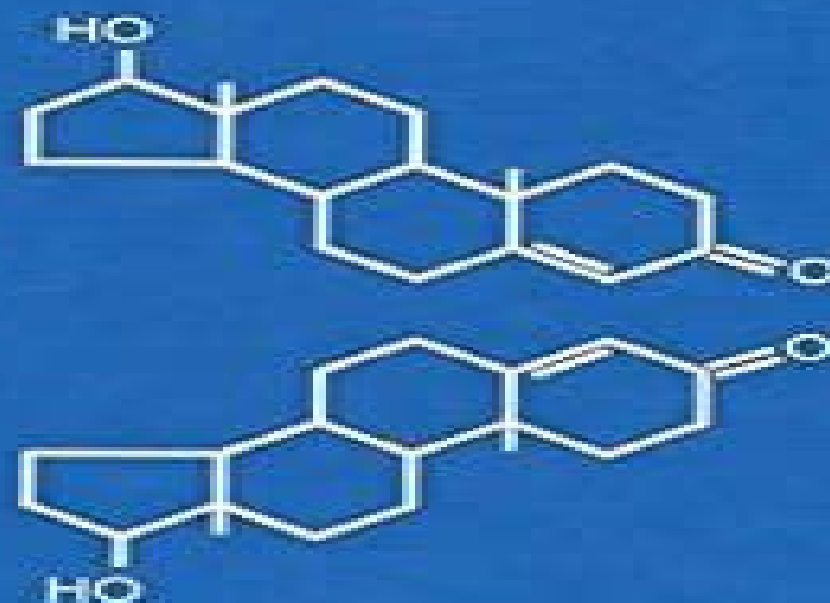
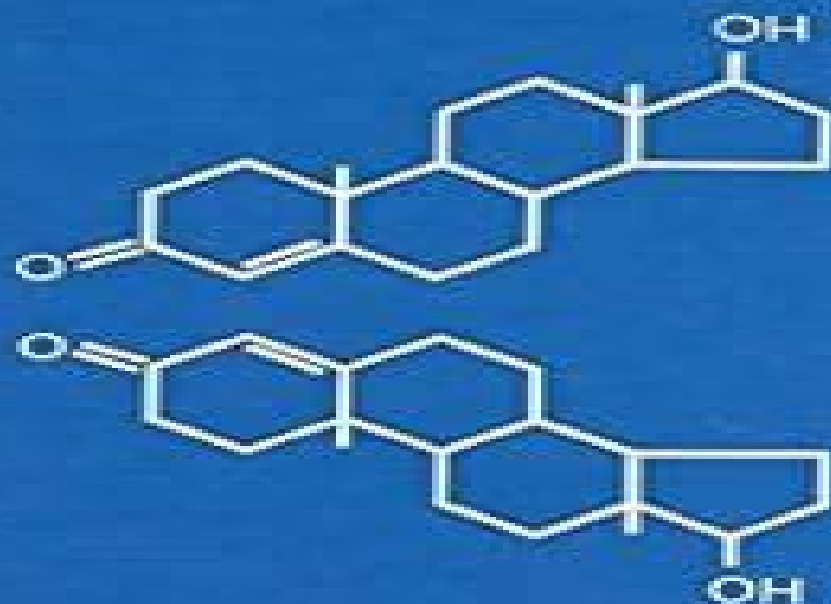


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web answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c explanation option c is incorrect because 566 kJ mol is the  $\Delta h$  per mole of o<sub>2</sub> and not for 2 moles even though options b and d are

expressed in grams they are still correct because 56

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web  $q = 140 \text{ g} \times 4.184 \text{ J}/160^\circ\text{C} \times q = 9372.16 \text{ J}$   $35.5 \text{ g} \times 1 \text{ mole KCN} / 0.5453 \text{ moles}$   $65.1 \text{ g} / 0.5453 \text{ mole moles}$  5 when a 15 g sample of  $SnCl_2$  dissolves in 80 g of water in a calorimeter the temperature drops from  $42^\circ\text{C}$  to  $15^\circ\text{C}$  calculate the  $\Delta H$  for the process

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