



The Gun Review

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{K} + 3\text{Al}_2\text{O}_3 + 6\text{S}$
 $1\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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web expert answer transcribed image text balancing equations and simple stoichiometry balance the following equations
 $1\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{C}_6\text{H}_{10}\text{O}_2 + \text{O}_2 \rightarrow 2\text{CO}_2 + \text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} \rightarrow \text{KBr} + \text{CO}_2 + \text{GaBr}_3 + \text{NaBr} + \text{SnO}_4$
 $4\text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{N}_2 + 5\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $1\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} \rightarrow 1\text{KBr} + 1\text{CO}_2 + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{N}_2 + 6\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$ using the following equation $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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