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web donor whereas a base generically represented by B is a proton acceptor when acids and bases combine they swap a proton H^+ $B + H^+ \rightarrow HB$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then $\text{pH} = -\log 0.100 = 1.000$ pH log 0.100 = 1.000 b titrant volume 12.50 mL

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