

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Study Guide To Chemistry

John W. Hill



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Table of Contents Study Guide To Chemistry

1. Understanding the eBook Study Guide To Chemistry
 - The Rise of Digital Reading Study Guide To Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide To Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide To Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide To Chemistry
 - Personalized Recommendations
 - Study Guide To Chemistry User Reviews and Ratings
 - Study Guide To Chemistry and Bestseller Lists
5. Accessing Study Guide To Chemistry Free and Paid eBooks
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- Study Guide To Chemistry Budget-Friendly Options
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 - Study Guide To Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide To Chemistry
 - Highlighting and Note-Taking Study Guide To Chemistry
 - Interactive Elements Study Guide To Chemistry
- 8. Staying Engaged with Study Guide To Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide To Chemistry
- 9. Balancing eBooks and Physical Books Study Guide To Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide To Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide To Chemistry
 - Setting Reading Goals Study Guide To Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide To Chemistry
 - Fact-Checking eBook Content of Study Guide To Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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