# THE ULTIMATE GENERAL CHEMISTRY 1 STUDY GUIDE

- · Matter and Measurement
  - Matter: anything that has mass and occupies space
  - Properties of matter: physical properties and chemical properties
  - States of matter: solid, liquid, gas
  - Measurements: SI Units, significant figures, dimensional analysis
- Atoms, Molecules, and Ions
  - Atomic structure: protons, neutrons, and electrons
  - Periodic table and trends: electronegativity, ionization energy, electron affinity
  - > Chemical bonding: ionic bonding, covalent bonding, metallic bonding
  - Chemical formulas and equations: empirical and molecular formulas, balancing chemical equations
- Stoichiometry
  - > Limiting reactants and percent yield
  - Mole concept and Avogadro's number
  - > Molarity and molality
  - Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
- Chemical Reactions and Equations
  - Types of chemical reactions: synthesis, decomposition, single replacement, double replacement, combustion
  - Redox reactions and balancing redox equations
  - Reaction rates: factors affecting reaction rates, rate laws, and reaction mechanisms
- ◆ Gases
  - Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
  - Kinetic Molecular Theory of gases: pressure, temperature, volume, and the number of moles of gas particles
  - Natural gases: van der Waals equation

# **Study Guide For General Chemistry And College Chemistry**

**Martin Silberberg** 

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