

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Study Guide For Changs Chemistry

S Marginson



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Table of Contents Study Guide For Changs Chemistry

1. Understanding the eBook Study Guide For Changs Chemistry
 - The Rise of Digital Reading Study Guide For Changs Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Changs Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Changs Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Changs Chemistry
 - Personalized Recommendations
 - Study Guide For Changs Chemistry User Reviews and Ratings
 - Study Guide For Changs Chemistry and Bestseller Lists
5. Accessing Study Guide For Changs Chemistry Free and Paid eBooks
 - Study Guide For Changs Chemistry Public Domain eBooks
 - Study Guide For Changs Chemistry eBook Subscription Services
 - Study Guide For Changs Chemistry Budget-Friendly Options
6. Navigating Study Guide For Changs Chemistry eBook Formats

- ePub, PDF, MOBI, and More
- Study Guide For Changs Chemistry Compatibility with Devices
- Study Guide For Changs Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide For Changs Chemistry
 - Highlighting and Note-Taking Study Guide For Changs Chemistry
 - Interactive Elements Study Guide For Changs Chemistry
- 8. Staying Engaged with Study Guide For Changs Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Changs Chemistry
- 9. Balancing eBooks and Physical Books Study Guide For Changs Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Changs Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Changs Chemistry
 - Setting Reading Goals Study Guide For Changs Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Changs Chemistry
 - Fact-Checking eBook Content of Study Guide For Changs Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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