



2. A right-angled triangle is shown.

Find the length of the hypotenuse.

### Section 1: Right-angled triangles

1. A right-angled triangle is shown.

Find the length of the hypotenuse.

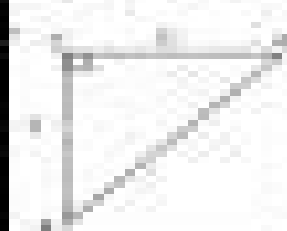
A right-angled triangle is shown. The base is 3 cm and the height is 4 cm. Find the length of the hypotenuse.



2. A right-angled triangle is shown. Find the length of the hypotenuse.



3. A right-angled triangle is shown. Find the length of the hypotenuse.



4. A right-angled triangle is shown. Find the length of the hypotenuse.

- (a) 5 cm
- (b) 6 cm
- (c) 7 cm
- (d) 8 cm
- (e) 9 cm
- (f) 10 cm

# Ssm Contemporary Trig

**Roy Dubisch**



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## **Table of Contents Ssm Contemporary Trig**

1. Understanding the eBook Ssm Contemporary Trig
  - The Rise of Digital Reading Ssm Contemporary Trig
  - Advantages of eBooks Over Traditional Books
2. Identifying Ssm Contemporary Trig
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ssm Contemporary Trig
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ssm Contemporary Trig
  - Personalized Recommendations
  - Ssm Contemporary Trig User Reviews and Ratings
  - Ssm Contemporary Trig and Bestseller Lists
5. Accessing Ssm Contemporary Trig Free and Paid eBooks
  - Ssm Contemporary Trig Public Domain eBooks
  - Ssm Contemporary Trig eBook Subscription Services
  - Ssm Contemporary Trig Budget-Friendly Options
6. Navigating Ssm Contemporary Trig eBook Formats

- ePub, PDF, MOBI, and More
- Ssm Contemporary Trig Compatibility with Devices
- Ssm Contemporary Trig Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ssm Contemporary Trig
  - Highlighting and Note-Taking Ssm Contemporary Trig
  - Interactive Elements Ssm Contemporary Trig
- 8. Staying Engaged with Ssm Contemporary Trig
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ssm Contemporary Trig
- 9. Balancing eBooks and Physical Books Ssm Contemporary Trig
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ssm Contemporary Trig
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ssm Contemporary Trig
  - Setting Reading Goals Ssm Contemporary Trig
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ssm Contemporary Trig
  - Fact-Checking eBook Content of Ssm Contemporary Trig
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely becuase bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture  $\text{HNO}_3$  3 mL  $\text{H}_2\text{SO}_4$  4 mL in the 125 mL flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 mL erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 Å bond distances this will result in an optimization failure as the atoms are even closer for the programme to accommodate their electrons demonstration of creating the p arenium cation by the nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion  $\text{NO}_2^+$  which acts as the electrophilic species

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and p bromonitrobenzene

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