

# Speaking the language of desire

THE FILMS OF  
CARL  
DREYER

By  
Raymond Carney



# Speaking The Language Of Desire The Films Of Carl Dreyer

**J Ma**



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monoxide co determine the mass of iodine  $\text{I}_2$  which could be produced b if in the above situation only 0 160 moles of iodine  $\text{I}_2$   
was produced i what mass of iodine was produced

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mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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 $\text{Cu} + 2 \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$  text cu 2 text agno 3 rightharpoon 2 text ag text cu no 3 text 2 cu 2 agno3 2 ag cu no3 2

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the balanced equation shown below if the reaction of 90 6 grams of co produces 36 7 grams of  $\text{C}_3\text{H}_8$  what is the percent  
yield  $3 \text{CO} + 7 \text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3 \text{H}_2\text{O}$

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web start with 8 3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8 3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149 4 \text{ g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137 1 \text{ g H}_2\text{O}}{149 4 \text{ g H}_2\text{O}} \times 100$

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web 7 Na + 2  $\text{C}_2\text{O}_4$  is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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