Henderson Hasselbalch Equation

Dissociation of an Acid

$$HA + H_2O \rightleftharpoons A^{\ominus} + H_3O^{\oplus}$$
Acid Water Conjugate Hydronium
Base

Equation

$$pH = pK_a + log_{10} \frac{[A^{\ominus}]}{[HA]}$$

pH:pH value of the solution

K : Dissociation constant of the acid

 $[A^{\Theta}]$: Concentration of the conjugate base (A^{Θ})

[HA]: Concentration of the acid (HA)

The Henderson Equation

Daniela Niemeyer

The Henderson Equation:

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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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