

# INORGANIC & ORGANIC CHEM LEC

Ms. Jazheerine Portugal | Chemical Equilibrium

## CHEMICAL EQUILIBRIUM

**Chemical equilibrium** is achieved when the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant.



closed system in equilibrium with its reactants in a closed system at their temperature



$\text{N}_2\text{O}_4$  and  $2\text{NO}_2$  gases at equilibrium



## EQUILIBRIUM CONSTANT

Table shows some experimental data for the reaction just described at 25°C. The gas concentrations are expressed in molarity, which can be calculated from the number of moles of gasses present initially and at equilibrium and the volume of the flask in liters.



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]} = 4.63 \times 10^{-3}$$

| Initial Concentrations (M) |                          | Equilibrium Concentrations (M) |                          | Ratio of Concentrations at Equilibrium         |                                                  |
|----------------------------|--------------------------|--------------------------------|--------------------------|------------------------------------------------|--------------------------------------------------|
| $[\text{NO}_2]$            | $[\text{N}_2\text{O}_4]$ | $[\text{NO}_2]$                | $[\text{N}_2\text{O}_4]$ | $\frac{[\text{NO}_2]}{[\text{N}_2\text{O}_4]}$ | $\frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$ |
| 0.000                      | 0.473                    | 0.0347                         | 0.443                    | 0.0781                                         | $4.63 \times 10^{-3}$                            |
| 0.0093                     | 0.464                    | 0.0497                         | 0.404                    | 0.123                                          | $4.64 \times 10^{-3}$                            |
| 0.0193                     | 0.450                    | 0.0475                         | 0.391                    | 0.122                                          | $4.60 \times 10^{-3}$                            |
| 0.0400                     | 0.410                    | 0.0923                         | 0.394                    | 0.234                                          | $4.60 \times 10^{-3}$                            |
| 0.269                      | 0.090                    | 0.0784                         | 0.0774                   | 0.277                                          | $4.68 \times 10^{-3}$                            |

We can generalize this phenomenon with the following reaction at equilibrium:



For a reversible reaction at equilibrium and a constant temperature, a certain ratio of reactant and product concentrations has a constant value,  $K$  (the **equilibrium constant**).

$$K = \frac{[\text{C}]^c[\text{D}]^d}{[\text{A}]^a[\text{B}]^b}$$

## HOMOGENEOUS EQUILIBRIA

**Homogeneous equilibrium** applies to reactions in which all reacting species are in the same phase.



$$K_c = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$K_P = \frac{P_{\text{NO}_2}^2}{P_{\text{N}_2\text{O}_4}}$$

**Sample Questions:**

- (a)  $\text{HF}(\text{aq}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{F}^-(\text{aq})$   
 (b)  $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$   
 (c)  $\text{CH}_3\text{COOH}(\text{aq}) + \text{C}_2\text{H}_5\text{OH}(\text{aq}) \rightleftharpoons \text{CH}_3\text{COOC}_2\text{H}_5(\text{aq})$

(a)

$$K_c = \frac{[\text{H}_3\text{O}^+][\text{F}^-]}{[\text{HF}]}$$

(b)

$$K_c = \frac{[\text{NO}_2]^2}{[\text{NO}]^2[\text{O}_2]} \quad K_P = \frac{P_{\text{NO}_2}^2}{P_{\text{NO}}^2 P_{\text{O}_2}}$$

# Ssm T A Des Inorganic Chem 4e

**Francesca Deganello, Enrico  
Traversa, Ana C. Tavares**



## Ssm T A Des Inorganic Chem 4e:

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*Mechanisms of Inorganic and Organometallic Reactions*, 1983 **Canadian Journal of Chemistry**, 1986 *Nuclear Science Abstracts*, 1974 **Chemistry, a Sustainable Bridge from Waste to Materials for Energy and Environment** Francesca Deganello, Enrico Traversa, Ana C. Tavares, 2021-03-08 **Photocatalytic Coatings for Air-Purifying, Self-Cleaning and Antimicrobial Properties** Anibal Maury-Ramirez, 2018-10-02 This book is a printed edition of the Special Issue Photocatalytic Coatings for Air Purifying Self Cleaning and Antimicrobial Properties that was published in *Coatings Bulletin signalétique*, 1971 *Cumulated Index Medicus*, 1982 *Energy Research Abstracts*, 1990 Semiannual with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works deriving from DOE other related government sponsored information and foreign nonnuclear information Arranged under 39 categories e g Biomedical sciences basic studies Biomedical sciences applied studies Health and safety and Fusion energy Entry gives bibliographical information and abstract Corporate author subject report number indexes

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