Structure of Inorganic Solids

Inorganic solids can be built of discreet molecules; however, most consist of infinite arrays of atoms or ions. Half of the elements are metals in which close-packed atoms are held together by delocalised electrons. The noble gases form solids in which close-packed atoms are held together by dispersion forces. Other elements consist of infinite networks of more localised bonds, Ionic solids form infinite networks held together by coulombic forces between ions of opposite charge - these structures are the basis of crystallinity.

Close packing in solids



Close packing of spheres in 1 dimension





Close packing in 2 dimensions

(at co-ordination number = 4). (b) co-ordination number = 6)

The best close packed arrangement is b - all sphere are in contact and 6 surround the central sphere. However, for ions, a is preferred in an open structure, where 9 spheres surround the central sphere, only 4 are in contact. B has the same charges in contact, which provides a high repulsive energy, where a has a much lower coulombic energy.

In 3D, there are two ways in which the above and below layers can pack. In both, the central sphere will be surrounded by 3 above and 3 below. making a co-ordination of 12.

Hexagonal close packing

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B Pattern repeats after

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Cubic close packing

ph. B C A.

Pattern repeats after evenly three rows.





Hexagonal close packing



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In hcp the spheres fill 74% of the available space - 26% present as holes in the structure

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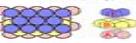
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Francis S. Galasso

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