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# The Hero Thorndike Press Large Print Juvenile Series Large Print

**RJ Shavelson**



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web log  $[\text{h}^+] = 10^{-4.05} = 8.91 \times 10^{-5} \text{ mol dm}^{-3}$  log pka ph 3.75  $4.05 - 0.30 = 10^{-0.30} = 0.50$  and  $\text{hcooh}$  0.1 so  $\text{hcoo}^-$  0.200 mol dm<sup>-3</sup> amount vol x

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