

Physical Science Study Guide Unit 2- Atomic Structure

1) What are the three subatomic particles, and where are they located in an atom?

Answer: protons and neutrons located within the nucleus, with electrons in orbitals surrounding the nucleus

2) Which subatomic particles are used to calculate the atomic number? Which subatomic particles are used to calculate the mass number?

Answer: the number of protons in an atom, protons and the number of neutrons

3) The different rows of the periodic table represent what about the electrons of an atom?

Answer: a new outer shell of electrons

4) What is a bond? What are atoms trying to do when they bond?

Answer: to attain a stable valence electron configuration (8 valence electrons) by bonding with other atoms. by forming covalent bonds

5) Where on the periodic table does it tell you the number of valence electrons an element has?

Answer: the column number that it is in

6) The atoms on the periodic table are organized by an increasing number of what?

Answer: number of protons

7) Why are the Noble gases unreactive?

Answer: Because they're reluctant to share electrons from their filled outer electron shells

8) Where on the periodic table do you find elements that are reactive because they have high electronegativity?

Answer: top right

9) Which elements are the best conductors: Metals or nonmetals?

Answer: Copper, Nickel, Tin

10) Where are nonmetals located on the periodic table?

Answer: right

11) Why is Hydrogen's Mass Number the same as its Atomic Number?

Answer: Presence of no neutron

12) Ionic bonds are most likely to occur when the number of valence electrons is close to what numbers?

Answer: The name of the compound is sodium fluoride with formula NaF.

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Mike Chapple, David Seidl



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