

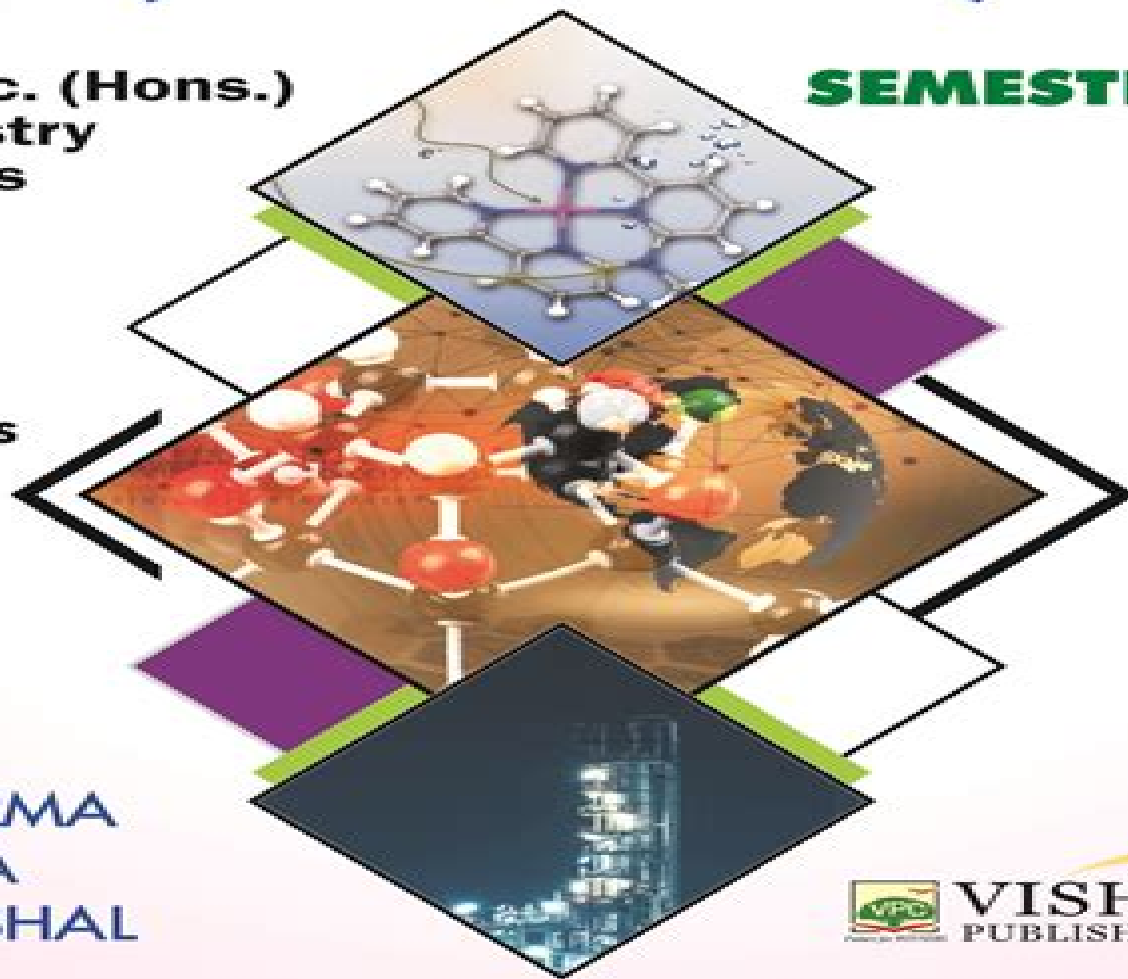
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Principles of Inorganic Chemistry Brian W. Pfennig, 2015-03-03 Aimed at senior undergraduates and first year graduate students this book offers a principles based approach to inorganic chemistry that unlike other texts uses chemical applications of group theory and molecular orbital theory throughout as an underlying framework This highly physical approach allows students to derive the greatest benefit of topics such as molecular orbital acid base theory band theory of solids and inorganic photochemistry to name a few Takes a principles based group and molecular orbital theory approach to inorganic chemistry The first inorganic chemistry textbook to provide a thorough treatment of group theory a topic usually relegated to only one or two chapters of texts giving it only a cursory overview Covers atomic and molecular term symbols symmetry coordinates in vibrational spectroscopy using the projection operator method polyatomic MO theory band theory and Tanabe Sugano diagrams Includes a heavy dose of group theory in the primary inorganic textbook most of the pedagogical benefits of integration and reinforcement of this material in the treatment of other topics such as frontier MO acid base theory band theory of solids inorganic photochemistry the Jahn Teller effect and Wade s rules are fully realized Very physical in nature compare to other textbooks in the field taking the time to go through mathematical derivations and to compare and contrast different theories of bonding in order to allow for a more rigorous treatment of their application to molecular structure bonding and spectroscopy Informal and engaging writing style worked examples throughout the text unanswered problems in every chapter contains a generous use of informative colorful illustrations

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Inorganic Chemistry James E. House, 2012-12-31 Inorganic Chemistry Second Edition provides essential information for students of inorganic chemistry or for chemists pursuing self study The presentation of topics is made with an effort to be clear and concise so that the book is portable and user friendly The text emphasizes fundamental principles including molecular structure acid base chemistry coordination chemistry ligand field theory and solid state chemistry It is organized into five major themes structure condensed phases solution chemistry main group and coordination compounds with several

chapters in each There is a logical progression from atomic structure to molecular structure to properties of substances based on molecular structures to behavior of solids etc The textbook contains a balance of topics in theoretical and descriptive chemistry For example the hard soft interaction principle is used to explain hydrogen bond strengths strengths of acids and bases stability of coordination compounds etc Discussion of elements begins with survey chapters focused on the main groups while later chapters cover the elements in greater detail Each chapter opens with narrative introductions and includes figures tables and end of chapter problem sets This new edition features new and improved illustrations including symmetry and 3D molecular orbital representations expanded coverage of spectroscopy instrumental techniques organometallic and bio inorganic chemistry and more in text worked out examples to encourage active learning and to prepare students for their exams This text is ideal for advanced undergraduate and graduate level students enrolled in the Inorganic Chemistry course This core course serves Chemistry and other science majors The book may also be suitable for biochemistry medicinal chemistry and other professionals who wish to learn more about this subject area Concise coverage maximizes student understanding and minimizes the inclusion of details students are unlikely to use Discussion of elements begins with survey chapters focused on the main groups while later chapters cover the elements in greater detail Each chapter opens with narrative introductions and includes figures tables and end of chapter problem sets

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and more advanced inorganic theories are presented in this one textbook while showing the relationships between the two

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thermodynamic origin Determination of binary formation constants by pH metry and spectrophotometry Chapter 3 Reaction Mechanism of Transition Metal Complexes I Inert and labile complexes Mechanisms for ligand replacement reactions Formation of complexes from aquo ions Ligand displacement reactions in octahedral complexes acid hydrolysis base hydrolysis Racemization of tris chelate complexes Electrophilic attack on ligands Chapter 4 Reaction Mechanism of Transition Metal Complexes II Mechanism of ligand displacement reactions in square planar complexes The trans effect Theories of trans effect Mechanism of electron transfer reactions types outer sphere electron transfer mechanism and inner sphere electron transfer mechanism Electron exchange Chapter 5 Isopoly and Heteropoly Acids and Salts Isopoly and Heteropoly acids and salts of Mo and W structures of isopoly and heteropoly anions Chapter 6 Crystal Structures Structures of some binary and ternary compounds such as fluorite antiferite rutile antirutile cristobalite layer lattices CdI_2 BiI_3 ReO_3 Mn_2O_3 corundum perovskite Ilmenite and Calcite Chapter 7 Metal Ligand Bonding Limitation of crystal field theory Molecular orbital theory octahedral tetrahedral or square planar complexes bonding and molecular orbital theory Chapter 8 Electronic Spectra of Transition Metal Complexes Spectroscopic ground states Correlation and spin orbit coupling in free ions for 1st series of transition metals Orgel and Tanabe Sugano diagrams for transition metal complexes d^1 d^9 states Calculation of Dq B and parameters Effect of distortion on the d orbital energy levels Structural evidence from electronic spectrum Jahn Teller effect Spectrochemical and nephelauxetic series Charge transfer spectra Electronic spectra of molecular addition compounds Chapter 9 Magnetic Properties of Transition Metal Complexes Elementary theory of magneto chemistry Gouy's method for determination of magnetic susceptibility Calculation of magnetic moments Magnetic properties of free ions Orbital contribution effect of ligand field Application of magneto chemistry in structure determination Magnetic exchange coupling and spin state cross over Chapter 10 Metal Clusters Structure and bonding in higher boranes Wade's rules Carboranes Metal carbonyl clusters low nuclearity carbonyl clusters Total electron count TEC Chapter 11 Metal Complexes Metal carbonyls structure and bonding Vibrational spectra of metal carbonyls for bonding and structure elucidation Important reactions of metal carbonyls Preparation bonding structure and important reactions of transition metal nitrosyl dinitrogen and dioxygen complexes Tertiary phosphine as ligand

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Table of Contents Textbook Of Inorganic Chemistry

1. Understanding the eBook Textbook Of Inorganic Chemistry
 - The Rise of Digital Reading Textbook Of Inorganic Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Textbook Of Inorganic Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Textbook Of Inorganic Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Textbook Of Inorganic Chemistry
 - Personalized Recommendations
 - Textbook Of Inorganic Chemistry User Reviews and Ratings
 - Textbook Of Inorganic Chemistry and Bestseller Lists
5. Accessing Textbook Of Inorganic Chemistry Free and Paid eBooks
 - Textbook Of Inorganic Chemistry Public Domain eBooks
 - Textbook Of Inorganic Chemistry eBook Subscription Services

- Textbook Of Inorganic Chemistry Budget-Friendly Options
- 6. Navigating Textbook Of Inorganic Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Textbook Of Inorganic Chemistry Compatibility with Devices
 - Textbook Of Inorganic Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Textbook Of Inorganic Chemistry
 - Highlighting and Note-Taking Textbook Of Inorganic Chemistry
 - Interactive Elements Textbook Of Inorganic Chemistry
- 8. Staying Engaged with Textbook Of Inorganic Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Textbook Of Inorganic Chemistry
- 9. Balancing eBooks and Physical Books Textbook Of Inorganic Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Textbook Of Inorganic Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Textbook Of Inorganic Chemistry
 - Setting Reading Goals Textbook Of Inorganic Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Textbook Of Inorganic Chemistry
 - Fact-Checking eBook Content of Textbook Of Inorganic Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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