

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to acid (LBP)



Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and indicators solution?

Acid-Base Indicators

Acids are proton donors, Bases are proton acceptors.



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets protonated, change of the base

and check for weak acids/bases of the acid.

to find the level of the hydrogens and more the electrons

in the other atom

Remember: -OH is a weak base, -OH- is a strong base

1. Acid: Donates and Proton: Donates, so: Proton Donor

What if there is no proton transfer?

Acid-Base Reactions

Acids: Donates its proton to base

Base: Donates its lone pair to acid



Acid: Donates its proton to base, Base: Donates its lone pair to acid

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Acid Strength and pKa

The Ka value of an acid (acid) and its conjugate base (A-) are related by:



The acid is a stronger acid than its conjugate base

To compare the acid, the pKa is used, which is the pKa of the acid

pKa = -logKa

$$\text{pKa} + \text{pKb} = 14$$

to find the stronger acid, the lower the pKa value (it's to compare the Ka)

Acid is a stronger acid than its conjugate base



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base (the base)

The stronger the acid, the more stable the conjugate base.



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Choosing a Base to Deprotonate a Given Compound

Suppose you want to deprotonate the following compound: CC(=O)O

Deprotonate means remove the proton (H+)

First, find the pKa value: CC(=O)O pKa = 4.75

Find the chemical equation of the compound reacting with a conjugate base (B-)



The pKa of the conjugate acid must be higher than 10, so the reaction will be favored

Find a conjugate base with pKa > 10. Acetate is not good (pKa = 4.75)



Choosing an Acid to Protonate a Given Compound

Suppose you want to protonate the following compound: CC(=O)[O-]

Protonate means add the proton (H+)

To do that, find the chemical equation of the reaction with an acid (HA)



The pKa of the acid must be lower than 4.75, so the reaction will be favored

Find an acid with pKa < 4.75. Acetic acid is not good (pKa = 4.75)

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Acid Strength

What does it mean "strong acid"?

Strongly donating the proton to the base, so the acid is strong

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Morris Hein, Peter C. Scott



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