

CHEMISTRY



Periodic Table of Elements

- Elemental forms:**
 - Monatomic:** Noble gases; halogens, atoms, organ, beryllium, sodium, and sodium, beryllium.
 - Diatoms:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; liquid: bromine; solid: iodine.
 - Molecular solids:** sulfur.
 - Covalent solids:** Carbon (diamond, graphite).
 - Metallic:** Most common type of element; high coordination numbers; very dense.
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options.
 - Binary:** Elements exchange.
 - Complex:** Mixed elements.
 - Polymer:** Repeating units.
 - Non-molecular:** Crystal lattice.
 - Transition metal compounds:**
 - Coordination compounds:** Non-bonding ligands.
 - Organometallics:** Metal-C bond.
 - Complex:** metallic bond.
- Periodicity:** The chemical properties of elements repeat on periodic table.
 - Family of elements:** Columns in periodic table.
 - The addition of protons produces an expanded series.**

Periodic Trends

- Chemistry determined by the behavior of valence electrons.
- Atoms or ions get larger as you go down the column; electron shells are added.
- Cation and anion (same charge) decrease in size as the atomic number (Z) increases.

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to remove; electron removing becomes difficult as atomic energy.
- Atomic and ionic radii:** Size and properties; atomic radii increase down the table; ionic radii based on ionic charge and coordination.
- Electron affinity:** Depends on charge, atomic size, and coordination (typical type).
- Atomic weight:** Based on atomic mass data or molecular weight standard data.

Development of the Periodic Table

- Mixing elements more predicted using periodicity.
- Columns 1: Alkali metals.
- Columns 2: Alkaline earth metals.
- First transition series: Titanium-Copper.
- Columns 13: Boron group; fluorine, chlorine, bromine, iodine, and astatine.
- Columns 14: Heavy gases; helium, neon, argon, krypton, xenon, and radon.
- Actinides: Z = 90 or above; radioactive.
- Lanthanides: f-orbital electrons; series with Z = 58.



Proton and electron

The addition of protons produces an expanded series.

Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an atom of charge +q, Z = q electrons; for a cation of charge +q, Z = q electrons.
- Mass number (A):** A = Z + N, where N = the number of neutrons in the nucleus.
- Isotopes:** Atoms with the same Z, different A; an element can have one or more isotopes; any sample contains a number of isotopes; no practical work, use the average mass number in atomic weight (weighted average of natural isotopes for a given element).



Atomic Quantum Numbers & Orbitals

- Principal quantum number (n):** Describes the size of the orbital and the energy level. n = 1, 2, 3, ...
- Angular momentum quantum number (l):** Describes the shape of the orbital. l = 0, 1, 2, ... n-1.
- Magnetic quantum number (m):** Describes the orientation of the orbital. m = -l, -(l-1), ..., 0, ..., (l-1), l.
- Spin quantum number (s):** Describes the spin of the electron. s = +1/2 or -1/2.
- Orbitals:** Regions of space where the probability of finding an electron is high.

Many-Electron Atoms

Orbitals and spin are determined for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that orbitals have different energies; the filling of levels is guided by the Aufbau principle.



n	l	m _l	m _s	Orbitals
1	0	0	+1/2	1s
1	0	0	-1/2	1s
2	0	0	+1/2	2s
2	0	0	-1/2	2s
2	1	-1	+1/2	2p _x
2	1	-1	-1/2	2p _x
2	1	0	+1/2	2p _y
2	1	0	-1/2	2p _y
2	1	1	+1/2	2p _z
2	1	1	-1/2	2p _z

- Electron configuration:** Atomic orbital occupancy for an atom.
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons, one with spin up, m_s = +1/2, and one with spin down, m_s = -1/2.
- Hund's rule:** Electrons fill p, d, and f orbitals in a manner to maximize the number of unpaired electrons (half-filled orbitals, spin is maximized).
- Ionization potential (IP):** Energy required to remove electrons from an atom or ion.
 - First IP:** Removal of first valence electron.
 - Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 4 for most gas.

Students Guide To Chemistry

SJ Ball



Students Guide To Chemistry:

Unveiling the Magic of Words: A Overview of "**Students Guide To Chemistry**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Students Guide To Chemistry**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book is central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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